

# Question Player

Last Modified on 10/29/2024 10:42 am EDT

In Smartwork, you access the assignments created by your instructor through the Smartwork Question Player. This page contains an overview of the Question Player's features.

Hide All Answers

## Smartwork Assignments

### Assignment Preview Page

When you enter the assignment, you will first be taken to the Assignment Preview page, which gives you information about the assignment itself, as well as information about each of the questions in the assignment.

You can exit the Assignment Preview page at any time by clicking "Return to Activity List" in the top left-hand corner of the screen.

Smartwork 5 CHEMISTRY: AN ATOMS-FOCUSED APPROACH, 3E student@wnorton.edu

### Chapter 1

Welcome to Smartwork! This assignment is designed with rich feedback to guide you as you learn.

Due Date: 10/30/24 00:40:00

**i** This assignment has a time limit of 40 min. Grades are accepted until October 30th, 2024, at 11:59 PM (Eastern Time).

0 OF 10 QUESTIONS COMPLETED **BEGIN ASSIGNMENT**

Question	Type	Points	Attempt	Status
01 When 6.00 g of sulfur are combined with 6.00 g o...	Numeric Entry	- / 1	- / ∞	Not Started
02 This Smartwork course uses a fourstep approach ...	Numeric Entry	- / 1	- / ∞	Not Started
03 Use the particulate representations of matter A t...	Mixed	- / 3	- / ∞	Not Started
04 Use the representations shown below to determi...	Sorting	- / 1	- / ∞	Not Started

### Information About the Assignment

At the top part of the Assignment Preview page, you will see information about the assignment, as well as information about your progress in the assignment.

## Chapter 1

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04 Use the representations shown below to determi...	Sorting	- / 1	- / ∞	Not Started

You will see the following information:

- The title of the assignment
- A brief description of the assignment, provided by your instructor
- Whether or not the assignment is graded or not
- A progress bar where you can keep track of how many questions you completed in the assignment
- Your grade for the assignment

### Assignment Table

Question	Type	Points	Attempt	Status
01 When 6.00 g of sulfur are combined with 6.00 g o...	Numeric Entry	- / 1	- / ∞	Not Started
02 This Smartwork course uses a fourstep approach ...	Numeric Entry	- / 1	- / ∞	Not Started
03 Use the particulate representations of matter A t...	Mixed	- / 3	- / ∞	Not Started
04 Use the representations shown below to determi...	Sorting	- / 1	- / ∞	Not Started
05 Mixtures can be separated into their component ...	Sorting	- / 1	- / ∞	Not Started
06 Given the diagram, answer the following questio...	Mixed	- / 3	- / ∞	Not Started
07 Watch the ChemTour animation below on intern...	Sorting	- / 2	- / ∞	Not Started
08 Rank these molecules in order from fewest atom...	Ranking	- / 2	- / ∞	Not Started
09 Two common materials we encounter in our daily...	Numeric Entry	- / 2	- / ∞	Not Started
10 Watch this animation on Precision and Accuracy, ...	Multiple Select	- / 3	- / ∞	Not Started

You will also see the Assignment Table, which tells you the following information about each question:

**Question:** In this column, you see each question's title, which gives you a sense of what topics the question will cover, as well as what the question will ask you to do.

**Type:** In this column, you see the type of question you will be expected to answer. The types of questions in Smartwork are as follows:

- Symbolic (Math) Equation
- Dropdown List

- **Chemical Equation**
- **Labeling**
- **Molecular Drawing**
- **Multiple Choice**
- **Multiple Select**
- **Numeric Entry**
- **Open Response**
- **Ranking**
- **Short Answer**
- **Sorting**
- **Mixed:** This means the question contains more than one question type.

For an in-depth description of each module type, please refer to the “Working on Questions” section of these help notes.

**Points:** In this column, you will see the number of points each question is worth. Your instructor determines the number of points each question is worth. The points you have for each question are displayed in the form of a fraction: the denominator displays the total number of points you can earn for that question, while the numerator displays the number of points that you have earned. If you have not attempted the question, you will see a dash in the numerator.

**Attempt:** In this column, you will see the number of attempts you have for each question. Your instructor sets the number of attempts you have for each question. Like the points section, the attempts you have for each section are displayed in the form of a fraction: the denominator displays the total number of attempts you have for that question, while the numerator displays the **number of attempts you have used**. Sometimes in the denominator you will see an infinity symbol,  $\infty$ . This means that your instructor gave you an unlimited number of attempts for this question.

**Status:** In this column, you will see the status of the problem. The possible statuses for each problem are:

*Not started*– you have not started the problem.

*No attempts*– you have entered into the problem, but have not made any attempts.

*In progress*– you have entered into the problem and made an attempt on it.

*Completed*– you have either exhausted your attempts, gotten the answer correct, or have given up and viewed the Solution for this question.

## Entering into an Assignment

To enter into an assignment, do one of the following:

1. To enter an assignment for the first time, click on the button on top of the Assignment Table that says “Begin Assignment.” This will take you to the first question in the assignment.

## Chapter 1

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**Note:** If you are *returning* to the assignment, this button will change to say “Resume Assignment.” It will take you to the question you were last working on. If you are returning to the assignment after you have submitted your grade, either to review your work or to work the questions again for practice, this button will say “Review Assignment.” It will take you to the first question in the assignment.

2. To enter directly into a particular question, click on the arrow button next to the question you would like to enter. Once in the assignment, you will be able to move between questions.

## Chapter 1

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### Navigating between the Assignment List and individual assignments

Note that when you click on an assignment from the main assignment list, it will open up in a new tab with a unique URL. If you are navigating between your main Direct Landing Page for the course, the assignment list, and individual assignments, remember to note which tab you are on. If you want to save the DLP as a bookmark, make sure you're saving that main course page and not that of an individual assignment, for ease of future use!

# Working on Questions

## Question Interface

In addition to giving you a space to answer the question, the Question Interface gives you information about the question, as well as information on how the question fits into the overall assignment.

### Information about the Assignment

The screenshot displays the Question Interface. At the top, a black bar contains navigation and status information: a back arrow, 'Chapter 1', 'Assignment Due Date', 'Due Date: 10/30/24', a 'Grade' section with a bar chart and '-- %', a timer showing '00:39:35' and 'Time Limit', and a user profile 'student@wwnorton.edu' with a settings gear. Below this bar, a red link says 'Click here to return to the Assignment Preview Page'. A red information icon is followed by the text: 'This is a Numeric Entry question / It is worth 1 point / You have unlimited attempts / There is no attempt penalty'. The main content area shows '01 Question (1 point)' with a link to 'Open Ebook section 1.1'. The question text reads: 'When 6.00 g of sulfur are combined with 6.00 g of oxygen, 12.00 g of sulfur dioxide (SO<sub>2</sub>) are formed.' Below this, a '1st attempt' section is visible. The question asks: 'What mass of oxygen would be required to convert 5.00 g of sulfur into sulfur trioxide (SO<sub>3</sub>)?' with an input field and 'g' next to it. A 'See Periodic Table' link is also present. At the bottom, an 'Assignment Status Bar' shows '0 OF 10 QUESTIONS COMPLETED', '1 of 10 Questions NEXT >', and a 'SUBMIT ANSWER' button.

In the black bar at the top of the interface, you can see the following information pertaining to the assignment:

- The assignment due date, also known as the “Grades Accepted Until Date,” if your instructor set one
- Your grade for this assignment
- The time limit you have to complete the assignment, if your instructor set one
  - **Note that if your instructor set a time limit, the timer in this black bar will begin counting down the moment you enter into the assignment. There is no way to pause the timer once you enter into an assignment.**

The black bar also contains a back button, which allows you to return to the Assignment Preview page at any time.

Additionally, you can track your progress in the assignment via the Assignment Status Bar, located at the bottom of the interface. This tells you the number of questions you have successfully completed in the assignment.

### Information about the QuestionQ

Chapter 1 Due Date: 10/30/24 SCORE -- % 00:39:35 student@wnorton.edu

This is a Numeric Entry question / It is worth 1 point / You have unlimited attempts / There is no attempt penalty

**Question Information**

**01 Question** (1 point) [Open Ebook section 1.1](#)

When 6.00 g of sulfur are combined with 6.00 g of oxygen, 12.00 g of sulfur dioxide (SO<sub>2</sub>) are formed.

1st attempt

[See Periodic Table](#)

What mass of oxygen would be required to convert 5.00 g of sulfur into sulfur trioxide (SO<sub>3</sub>)?  
 g

0 OF 10 QUESTIONS COMPLETED 1 of 10 Questions NEXT ▶ [SUBMIT ANSWER](#)

Located underneath the black bar at the top of the interface, the Question Information Bar gives you the following information about the question you are answering:

- **Question Type**– this tells you what type of question you are answering (for example, whether it is a dropdown list question, a multiple choice question, a ranking question, etc). We will discuss the different types of questions in Smartwork in the next section.
- **Points**– This tells you the number of points this question is worth. You can also find the number of points a question is worth next to the word “Question” in the interface.
- **Attempts**– This tells you how many attempts you have to answer this question, as well as how many attempts you have used. For example, if the system says “You have 1 of 2 attempts remaining,” it means that you have two total attempts to answer the question, and that you have already used one of these attempts.
- **Attempt Penalty**– This tells you the percentage of points you will get off your score for this question each time you submit an incorrect answer attempt. For example, if this question is worth 1 point, and there is a 10% attempt penalty, the maximum number of points you can earn if you get your first attempt incorrect is .90 points.

In addition to the information in the Question Information Bar, you can see the number of points the question is worth, as well as the question’s order in the assignment, next to the word “Question.”

**01 Question** (1 point)

[See page 5](#)

## Introductions

Some questions may contain an introduction. Introductions may consist of a short paragraph, video, or image with information you may need to answer the question. Introductions appear underneath the question number and the number of points the question is worth.

This is a Numeric Entry question / It is worth 2 points / You have unlimited attempts / There is no attempt penalty

### 09 Question (2 points)

[Open Ebook section 1.7](#)

Two common materials we encounter in our daily lives are radiator coolant in automobiles and dry ice. The coolant in an automobile radiator, a solution of antifreeze (a glycol such as ethylene or propylene glycol) and water, freezes at  $-39\text{ }^{\circ}\text{C}$  and boils at  $110\text{ }^{\circ}\text{C}$ . The temperature of the dry ice (solid carbon dioxide) used in some ice cream vending carts is  $-78\text{ }^{\circ}\text{C}$ .

#### 1st attempt

##### Part 1 (1 point)

[See Periodic Table](#)

Convert the boiling point of radiator coolant ( $110\text{ }^{\circ}\text{C}$ ) to degrees Fahrenheit.   $^{\circ}\text{F}$

##### Part 2 (1 point)

Convert the temperature of dry ice ( $-78\text{ }^{\circ}\text{C}$ ) to the Kelvin scale.  K

0 OF 10 QUESTIONS COMPLETED

[PREVIOUS](#) 9 of 10 Questions [NEXT](#)

[SUBMIT ANSWER](#)

### Hints

If your instructor has enabled hints, you will be able to see them by clicking on the “See Hint” button, located to the right of the question stem. The icon associated with this is a light bulb. Clicking on this button will open up the Hint in a pop-up window.

### 01 Question (2 points)

#### 1st attempt

[See Hint](#)

In  $84.0\text{ g}$  of  $\text{NH}_4\text{NO}_3$ , how many grams of N are there?

9

### Ebook Reference

If your instructor has enabled you to see the ebook reference, you can find it in the upper right-hand corner of the question. If you have purchased ebook access, clicking on the ebook reference will open up the ebook in a separate tab, and direct you to the section in the text that goes over the concepts you need to solve the question.

Chapter 1 Due Date: 10/30/24 SCORE -- % 00:39:35 student@wnorton.edu

This is a Numeric Entry question / It is worth 1 point / You have unlimited attempts / There is no attempt penalty

**01 Question** (1 point) [Open Ebook section 1.1](#)

When 6.00 g of sulfur are combined with 6.00 g of oxygen, 12.00 g of sulfur dioxide (SO<sub>2</sub>) are formed.

1st attempt

See Periodic Table

What mass of oxygen would be required to convert 5.00 g of sulfur into sulfur trioxide (SO<sub>3</sub>)?  
 g

0 OF 10 QUESTIONS COMPLETED 1 of 10 Questions NEXT ▶ SUBMIT ANSWER

### Question Module Types

The question is located in the center of the interface and is composed of two main components: the question stem, which is the question itself, and the question module, which is the vehicle by which you answer the question. These modules correspond to the Question Type displayed on the Assignment Preview page.

**01 Question** (2 points)

1st attempt

See Hint

In 84.0 g of NH<sub>4</sub>NO<sub>3</sub>, how many grams of N are there?  
 g

- **Chemical Equation**– the question will ask you to balance or solve a chemical equation.
- **Symbolic (Math) Equation**– the question will ask you to solve a math equation.
- **Dropdown List**– the question will ask you to select the correct answer from a dropdown menu of options.
- **Labeling**– the question will ask you to label an image.
- **Molecular Drawing**– the question will either ask you to draw a molecule, or ask you to select part of a molecule for your answer.
- **Multiple Choice**– the question will ask you to select one answer option.
- **Multiple Select**– the question will ask you to select one or more answer options.
- **Numeric Entry**– the question will ask you to enter your answer in the form of a number. To solve this type of question, you may be asked to perform a calculation.
- **Ranking**– the question will ask you to rank words or images in a particular order.
- **Short Answer**– the question will ask you to give your answer in the form of a word or a short phrase.
- **Sorting**– the question will ask you to sort images or words into categories.



## Submitting Answers

The screenshot shows a question interface with a dark header bar. On the left, it says 'Chapter 1'. In the center, 'Due Date: 10/30/24' is displayed next to a 'SCORE' icon and a percentage sign. On the right, a timer shows '00:24:18' and a user profile icon with the email 'student@wnorton.edu'. Below the header, a red information icon is followed by the text: 'This is a Numeric Entry question / It is worth 2 points / You have unlimited attempts / There is no attempt penalty'. The main content area is titled '09 Question (2 points)' with a link to 'Open Ebook section 1.7'. The question text reads: 'Two common materials we encounter in our daily lives are radiator coolant in automobiles and dry ice. The coolant in an automobile radiator, a solution of antifreeze (a glycol such as ethylene or propylene glycol) and water, freezes at -39 °C and boils at 110 °C. The temperature of the dry ice (solid carbon dioxide) used in some ice cream vending carts is -78 °C.' Below the question, a '1st attempt' section is shown with a dropdown arrow and an up arrow. It contains two parts: 'Part 1 (1 point)' with a 'See Periodic Table' link and a prompt to convert 110 °C to Fahrenheit, with an input field containing '230'; and 'Part 2 (1 point)' with a prompt to convert -78 °C to Kelvin, with an input field containing '195'. At the bottom, a progress bar shows '0 OF 10 QUESTIONS COMPLETED', navigation buttons for 'PREVIOUS' and 'NEXT', and a 'SUBMIT ANSWER' button.

Once you have answered the question, you can submit your answer by clicking on the “Submit Answer” button in the bottom-right corner. Once you submit your answer, the system will notify you as to whether your answer is correct or incorrect. It will do this via a red X next to the module if the answer is incorrect, or a green checkmark next to the module if the answer is correct.

**If your instructor has enabled you to see feedback when answering questions,** you will also see feedback in a pop-up window. For an in-depth discussion about the different types of feedback, please refer to the “[Feedback](#)” section of these help notes.

You can submit answers until you run out of answer attempts. The system tells you which answer attempt you are on above the module (under the word “Question”).

In the Question Interface, your most recent answer attempt will stack above the previous attempt. Clicking on the “arrow” for the attempt in that stack will expand it, allowing you to view the answers you gave in previous attempts.

Chapter 1 Due Date: 10/30/24 SCORE 100% 00:20:55

**09 Question** (2 points) [Open Ebook section 1.7](#)

Two common materials we encounter in our daily lives are radiator coolant in automobiles and dry ice. The coolant in an automobile radiator, a solution of antifreeze (a glycol such as ethylene or propylene glycol) and water, freezes at  $-39\text{ }^{\circ}\text{C}$  and boils at  $110\text{ }^{\circ}\text{C}$ . The temperature of the dry ice (solid carbon dioxide) used in some ice cream vending carts is  $-78\text{ }^{\circ}\text{C}$ .

**Solution**

**Part 1 (1 point)** [Explanation](#) [See Hint](#)

Convert the boiling point of radiator coolant ( $110\text{ }^{\circ}\text{C}$ ) to degrees Fahrenheit.  °F

**Part 2 (1 point)** [Explanation](#) [See Hint](#)

Convert the temperature of dry ice ( $-78\text{ }^{\circ}\text{C}$ ) to the Kelvin scale.  K

**1st attempt**

Screenshot [PREVIOUS](#) 9 of 10 Questions [NEXT](#) [QUESTION COMPLETED](#)

Once you have exhausted your answer attempts, you give up on the question, or you get the answer correct, the “Submit Answer” button will deactivate, and instead will say “Question Completed.”

### Moving Between Questions

You can move between questions by clicking on the arrows at the bottom of the Question Editor.

Chapter 1 Due Date: 10/30/24 SCORE 100% 00:15:39

**09 Question** (2 points) [Open Ebook section 1.7](#)

Two common materials we encounter in our daily lives are radiator coolant in automobiles and dry ice. The coolant in an automobile radiator, a solution of antifreeze (a glycol such as ethylene or propylene glycol) and water, freezes at  $-39\text{ }^{\circ}\text{C}$  and boils at  $110\text{ }^{\circ}\text{C}$ . The temperature of the dry ice (solid carbon dioxide) used in some ice cream vending carts is  $-78\text{ }^{\circ}\text{C}$ .

**Solution**

**Part 1 (1 point)** [Explanation](#) [See Hint](#)

Convert the boiling point of radiator coolant ( $110\text{ }^{\circ}\text{C}$ ) to degrees Fahrenheit.  °F

**Part 2 (1 point)** [Explanation](#) [See Hint](#)

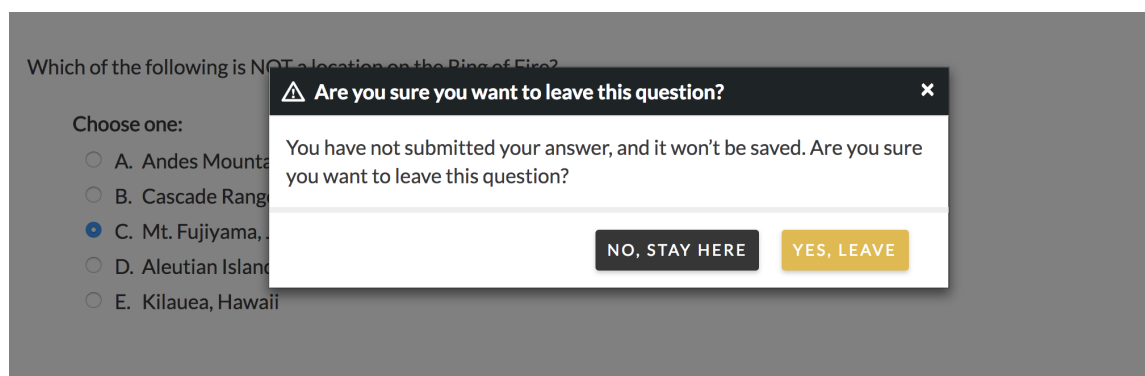
Convert the temperature of dry ice ( $-78\text{ }^{\circ}\text{C}$ ) to the Kelvin scale.  K

**1st attempt**

[PREVIOUS](#) 9 of 10 Questions [NEXT](#) [QUESTION COMPLETED](#)

If you began working on the question but did not submit your answer, you can still move on to another

question in the assignment. However, any work you have done on the question will not be saved. You will be reminded of this via a warning message in a modal window that reads, "You have not submitted your answer, and it won't be saved. Are you sure you want to leave this question?"



## Feedback

Depending on the settings your instructor set for the assignment, you may see two types of feedback: incorrect feedback if you get the answer incorrect, and correct feedback, if you get the answer correct. Additionally, your instructor may also allow you to view the Solution for a question, which is a detailed explanation of how to solve the problem.

### Viewing Feedback

**Your ability to view feedback and the Solution depends upon the settings your instructor sets for an assignment.**

You may be able to view **feedback** under the following circumstances, depending on your instructor's choices:

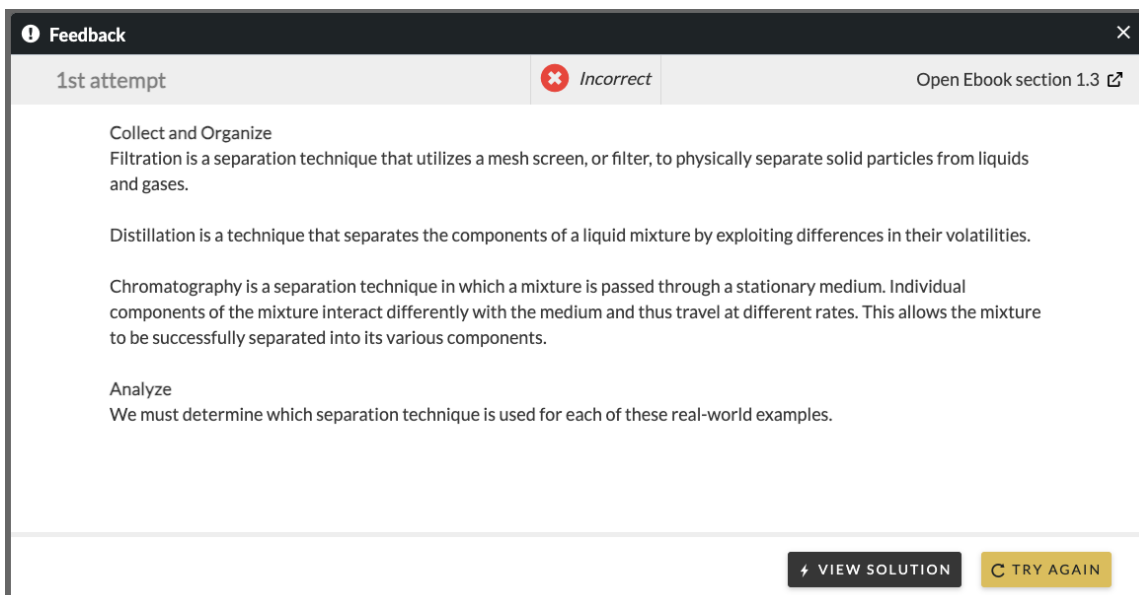
- After every answer attempt you give.
- After your final answer attempt (meaning, the last answer you give before you use up all of your answer attempts).
- After you get the answer correct.
- After the Grades Accepted Until date passes.

You may be able to view the **Solution** of a question under the following circumstances, depending on your instructor's choices:

- After you submit your first answer attempt, then you want to give up and you click "View Solution." (Note: if your instructor chooses this setting, viewing the Solution will mean you forfeit your remaining answer attempts.)
- After your final answer attempt or after you get the answer correct.
- Only after the Grades Accepted Until Date passes.

### Incorrect Feedback

Incorrect feedback triggers if you incorrectly answer the question. This feedback will be a few sentences explaining why you got the answer incorrect on a particular attempt. It will display in a separate pop-up window.



**Feedback** X

1st attempt Incorrect Open Ebook section 1.3 ↗

**Collect and Organize**  
Filtration is a separation technique that utilizes a mesh screen, or filter, to physically separate solid particles from liquids and gases.

Distillation is a technique that separates the components of a liquid mixture by exploiting differences in their volatilities.

Chromatography is a separation technique in which a mixture is passed through a stationary medium. Individual components of the mixture interact differently with the medium and thus travel at different rates. This allows the mixture to be successfully separated into its various components.

**Analyze**  
We must determine which separation technique is used for each of these real-world examples.

⚡ VIEW SOLUTION TRY AGAIN

In addition to displaying the incorrect feedback, the upper-left corner of the pop-up window also displays the answer attempt to which the feedback applies. Furthermore, if your instructor has allowed it, you will see the ebook page reference in the upper-right corner of the modal window. If you purchased ebook access, clicking on this link will take you to the section in the ebook that discusses the content in the question.

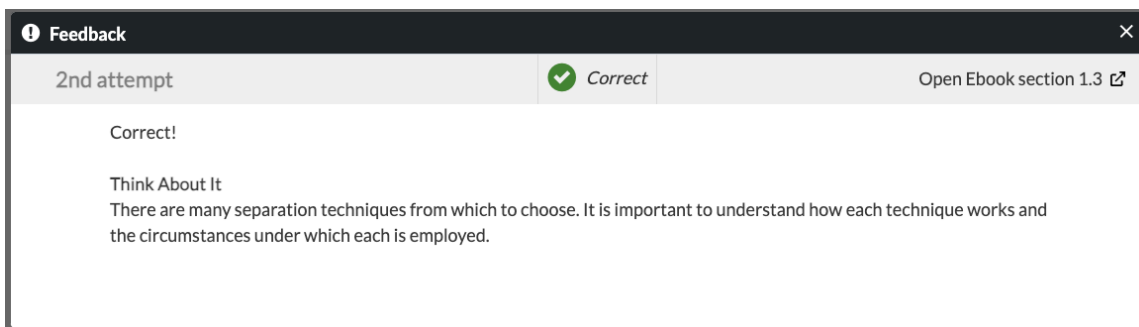
You can exit the incorrect feedback modal window by clicking on one of two buttons in the pop-up window's lower-right corner:

1. **View Solution**– Depending on the settings your instructor set, you may be able to view the Solution by clicking on the View Solution button. If you have remaining answer attempts, and you click the "View Solution" button, you will forfeit those attempts.
2. **Try Again**– Clicking the "Try Again" button will close the pop-up window and take you to a new question attempt. Your answer from the previous question attempt will be auto-populated in this new question attempt, so you can make adjustments to your previous answer based on the feedback. This option is only available to you if you have not exhausted your answer attempts.

You may also click the **X** in the upper right corner of the modal window, which will take you back to your answer attempt. In this case, you may click on the "View Solution" and "Try Again" buttons in the lower-right corner of the Question Interface (where the Submit Answer button was).

### Correct Feedback

Correct feedback triggers when you correctly answer a question. This feedback is a short paragraph explaining why you got the answer correct. This will display in a separate pop-up window.



In addition to displaying the correct feedback, the pop-up window also displays the attempt to which the feedback applies in the upper-left corner. If your instructor has enabled ebook references, you will be able to see it in the upper-right corner of the modal window. If you purchased ebook access, clicking on this link will take you to the section in the ebook that discusses the content in the question.

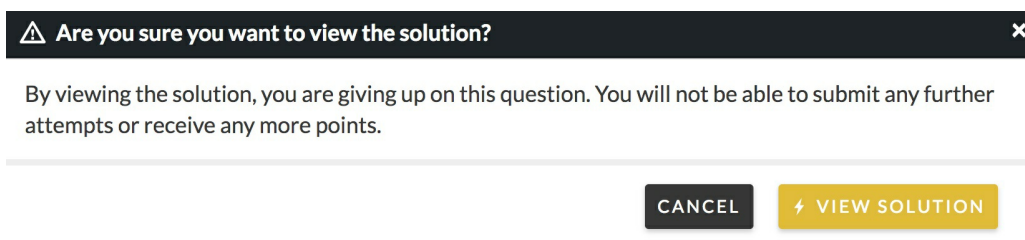
You can exit this pop-up window by clicking on the **X** in the upper-right corner.

### **Solution**

The Solution consists of the correct answer, as well as an in-depth explanation as to how to arrive at the correct answer. Like feedback, your ability to view the Solution depends on the settings your instructor set for the assignment. The Solution tab stacks above your last attempt in the Question Interface. The explanation of how to arrive at the Solution can be accessed by clicking the “Explanation” link in the Solution tab.

#### **If your instructor allows you to view the Solution at any time:**

- You can access the Solution **after you give your first incorrect attempt** from one of two places: the incorrect feedback modal, or the Question Interface.
- Once you click on this button, a message will pop up that reads, “By viewing the Solution, you are giving up on this question. You will not be able to submit any further attempts or receive any more points.”



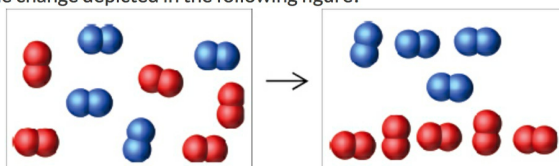
- The Solution will appear in the Question Interface, stacked above your last answer attempt. This gives you the correct answer.

▼ **Solution**

⚡ [Explanation](#)

💡 [See Hint](#)

How would you describe the change depicted in the following figure?



Choose one:

- A. A mixture of two gaseous elements is cooled to a temperature at which one of them condenses.
- B. A mixture of two gaseous compounds is heated to a temperature at which one of them decomposes.
- C. A mixture of two gaseous elements reacts together to form two compounds, one of which is a liquid.
- D. A mixture of two gaseous elements undergoes deposition.

- In the Solution tab, click on the **Explanation** link to see a detailed explanation of how to arrive at the correct answer. This explanation will appear in a pop-up window. If your instructor has enabled you to see the ebook reference, this will appear in the top-right corner of the window. You can exit the pop-up window by clicking on the **X** in the upper-right corner.

**If your instructor allows you to see the Solution after you get the answer correct or after you exhaust all of your attempts:**

- The Solution tab will automatically appear once you close the feedback window. There will not be a “View Solution” button. In the question interface, the Solution tab will be above your last attempt.

**09 Question** (2 points)

[Open Ebook section 1.7](#)

Two common materials we encounter in our daily lives are radiator coolant in automobiles and dry ice. The coolant in an automobile radiator, a solution of antifreeze (a glycol such as ethylene or propylene glycol) and water, freezes at  $-39\text{ }^{\circ}\text{C}$  and boils at  $110\text{ }^{\circ}\text{C}$ . The temperature of the dry ice (solid carbon dioxide) used in some ice cream vending carts is  $-78\text{ }^{\circ}\text{C}$ .

▼ **Solution**



**Part 1 (1 point)** ⚡ [Explanation](#)

Convert the boiling point of radiator coolant ( $110\text{ }^{\circ}\text{C}$ ) to degrees Fahrenheit.   $^{\circ}\text{F}$

**Part 2 (1 point)** ⚡ [Explanation](#)

Convert the temperature of dry ice ( $-78\text{ }^{\circ}\text{C}$ ) to the Kelvin scale.   $\text{K}$

> **1st attempt**



- You can access the Explanation by clicking on the "Explanation" link in the Solution tab.

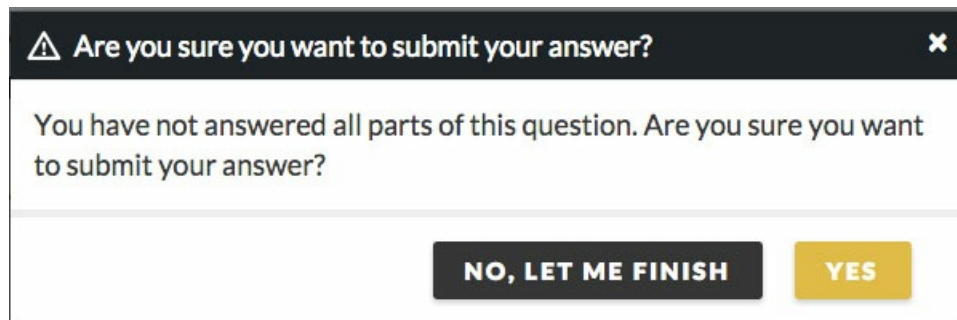
**If your instructor allows you to see the Solution after the Grades Accepted Until date passes:**

- You will not be able to access the Solution until after the Grades Accepted Until date passes. When you re-enter the assignment, you will see the Solution tab above your last attempt in the Question Interface.

**Multi-Part Questions**

Questions may contain more than one part. Each part may contain a different module type from the last. If enabled, each part may have its own Hint.

If you do not answer all parts of the question when you submit answer, you will encounter a pop-up message prompting you to answer all parts of the question.



Once you submit your answer, you may see feedback for each part in a feedback window, depending on the settings your instructor set for this assignment. If you exhaust your answer attempts, and have answered some but not all of the question parts correctly, you should get partial credit for your answers.

**Feedback**
×

1st attempt

✖ Part 1
✖ Part 2
✔ Part 3

Open Ebook section 1.4 [↗](#)

**✖ Part 1**

**Collect and Organize**  
Compare and contrast the definitions for solids, liquids, and gases.

**Analyze**  
Rationalize the phase represented by each frame for each molecule. Are we creating new molecules with a different atom-to-atom arrangement?

Recall the definitions for each of the terms associated with phase changes, as well as how the particulate model would change given the phase change.

---

**✖ Part 2**

**Analyze**  
Rationalize the phase represented by each frame for each molecule. Are we creating new molecules with a different atom-to-atom arrangement?

For the change to occur with water, for example, would we need to increase or decrease the temperature?

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**✔ Part 3**

**Think About It**  
This is a physical change. The atomic composition of each molecule is not changing.

When you view the Solution for a multi-part question, you will be able to view the Solution and Explanation for each part in the Solution tab.

▼ **Solution**
⚡

**Part 1 (1 point)** ⚡ [Explanation](#)

Which of the following outline the phase change(s) observed from the initial to final frame?

**Choose one or more:**

- condensation
- sublimation
- freezing
- melting
- vaporization
- deposition

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**Part 2 (1 point)** ⚡ [Explanation](#)

The temperature of the two frames is reported to be either 80°C or 135°C. What is the temperature of the right (products) based upon the observations?

°C

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**Part 3 (1 point)** ⚡ [Explanation](#)

The representations depict a  change.



## Assignments Reopened for Practice

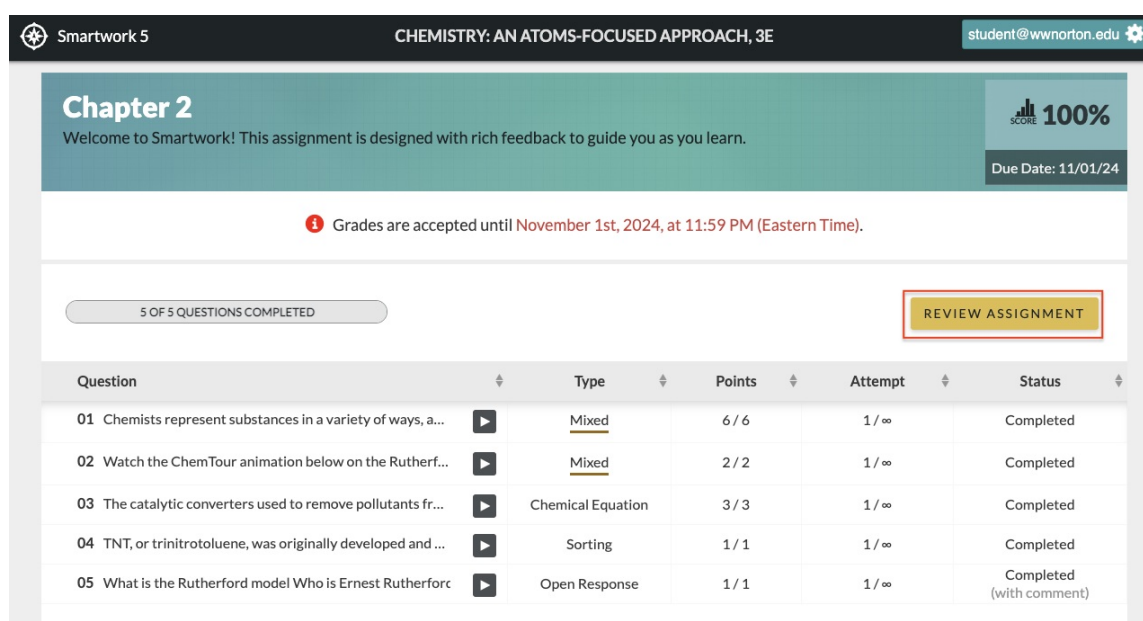
Depending on the settings your instructor set for the assignment, you may be able to return to an assignment after you submitted your grade and work on questions for ungraded practice.

You may do this under two circumstances: 1) any time after you complete a question, or 2) after the Grades Accepted Until date passes for the assignment.

If you do not have the option to work on questions for ungraded practice, you may still re-enter an assignment to review your work.

### Re-entering Assignments

To re-enter a completed assignment from the Assignment Preview page, click on the “Review Assignment” button in the upper-right corner. Or, click on the arrow next to the specific question you would like to review.



Smartwork 5 CHEMISTRY: AN ATOMS-FOCUSED APPROACH, 3E student@wnorton.edu

### Chapter 2

Welcome to Smartwork! This assignment is designed with rich feedback to guide you as you learn.

SCORE 100%

Due Date: 11/01/24

Grades are accepted until November 1st, 2024, at 11:59 PM (Eastern Time).

5 OF 5 QUESTIONS COMPLETED

REVIEW ASSIGNMENT

Question	Type	Points	Attempt	Status
01 Chemists represent substances in a variety of ways, a...	Mixed	6 / 6	1 / ∞	Completed
02 Watch the ChemTour animation below on the Rutherf...	Mixed	2 / 2	1 / ∞	Completed
03 The catalytic converters used to remove pollutants fr...	Chemical Equation	3 / 3	1 / ∞	Completed
04 TNT, or trinitrotoluene, was originally developed and ...	Sorting	1 / 1	1 / ∞	Completed
05 What is the Rutherford model? Who is Ernest Rutherfor...	Open Response	1 / 1	1 / ∞	Completed (with comment)

### Practicing Questions

To practice a completed question, do the following:

1. Open up the question you would like to practice. When you do this, the Solution and any previous attempts you made on the question will be collapsed in the Question Interface.
2. Click the “Practice” button in the bottom-right corner of the Question Interface.

This is a Mixed question / It is worth 6 points / You have unlimited attempts / There is no attempt penalty

### 01 Question (6 points)

[Open Ebook section 2.1](#)

Chemists represent substances in a variety of ways, and they need to be able to easily shift back and forth between these symbolic representations and names of the substances. They also need to be able to determine from any of these representations whether a pure substance is an element or a compound.

> Solution



> 1st attempt



5 OF 5 QUESTIONS COMPLETED

1 of 5 Questions NEXT ▶

↻ PRACTICE

✓ QUESTION COMPLETED

3. Re-work the question for ungraded practice.
  4. When done, click the “Check Practice” button in the bottom-right corner of the Question Interface. You will see if your practice attempt is correct or incorrect. If your instructor enabled feedback for the assignment, you will see feedback on your practice attempt.
  5. If desired, you may make another practice attempt. There is no limit to the number of practice attempts you can make.
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