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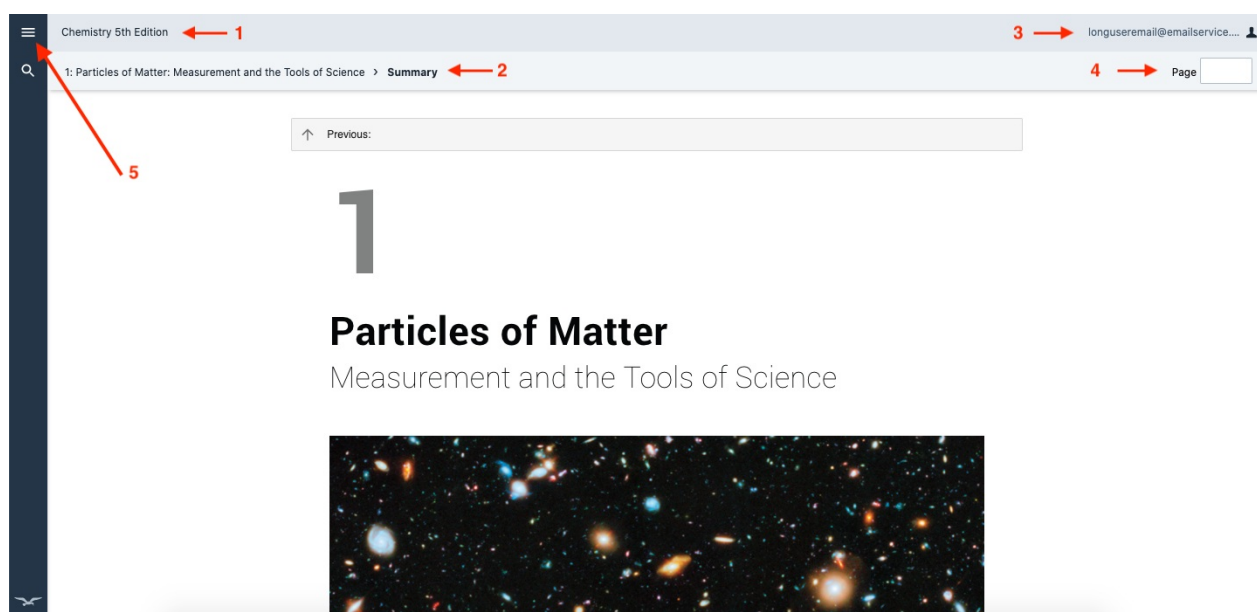
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4. Indicates what **Page** you are currently viewing.
5. Select the three horizontal lines to make the **Table of Contents** appear. More information about this feature can be found here.

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1: Particles of Matter: Measurement and the Tools of Science > **How and Why** Page

↑ Previous: 1: Particles of Matter: Measurement and the Tools of Science

1.1 How and Why

For thousands of years, humans have sought to better understand the world around us. For most of that time we resorted to mythological explanations of natural phenomena. Many once believed, for example, that the Sun rose in the east and set in the west because it was carried across the sky by a god driving a chariot propelled by winged horses.

In recent times we have been able to move beyond such fanciful accounts of natural phenomena to explanations based on observation and scientific reasoning. Unfortunately, this movement toward rational explanations has not always been smooth. Consider, for example, the contributions of a man whom Albert Einstein called the father of modern science, Galileo Galilei. At the dawn of the 17th century, Galileo used advanced telescopes of his own design to observe the movement of the planets and their moons. He concluded that they, like Earth, revolved around the Sun. However, this view conflicted with a belief held by many religious leaders of his time that Earth was the center of the universe. In 1633 a religious tribunal forced Galileo to disavow his conclusion that Earth orbited the Sun and banned him (or anyone) from publishing the results of studies that called into question the Earth-centered view of the universe. The ban was not completely lifted until 1835—nearly 200 years after Galileo's death.

In the last century, advances in the design and performance of telescopes have led to the astounding discovery that we live in an expanding universe that probably began 13.8 billion years ago with an enormous release of energy. In this chapter and in later ones, we examine some of the

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Sample Exercises 1.3, 1.4, 1.9

L08 Express uncertain values with the appropriate number of significant figures

Sample Exercise 1.5

L09 Distinguish between exact and uncertain values, evaluate the precision and accuracy of experimental results, and identify outliers

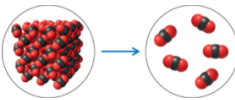
Sample Exercises 1.6, 1.7, 1.8

PARTICULATE PREVIEW

Matter and Energy

The temperature in outer space is 2.73 K. The temperature of dry ice (carbon dioxide, CO₂) is 70 times warmer, but still cold enough to keep ice cream frozen on a hot summer day. As you read [Chapter 1](#), look for ideas that will help you answer these questions:

- Particulate images of CO₂ as it sublimates are shown here. Which two phases of matter are involved in sublimation?
- What features of the images helped you decide which two phases were involved?
- What is the role of energy in this transformation of matter? Must energy be added or is energy produced?



↓ Next: How and Why

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Alloys and Medicine > Substitutional Alloys Page

CURRENT LOCATION

12: Solids: Crystals, Alloys, and Polymers


↳ Alloys and Medicine

↳ Substitutional Alloys

↑ Previous: Structures of Metals

12.3 Alloys and Medicine

The antibacterial properties of copper metal are attractive for coating surfaces in hospitals and in food service kitchens where an infection can prove deadly (Figure 12.11). However, pure copper has two disadvantages: it is both relatively soft and very malleable, which means that pure copper objects are easily bent and damaged. We can explain the malleability of Au, Cu, and other metals in terms of the relatively weak bonds between the atoms in their cubic closest-packed crystal structure. This arrangement gives the atoms in one layer the ability, under stress, to slip past atoms in an adjacent layer (Figure 12.12), but the overall crystal structure is still cubic closest-packed. The ease with which copper atoms slip past each other makes it easy to bend copper pipes used in plumbing, but it also makes them susceptible to damage. Additionally, copper reacts with air to produce blue-green copper hydroxides and carbonates.




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3: Stoichiometry: Mass, Formulas, and Reactions Page 82




↑ Previous: Questions and Problems

3

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Mass, Formulas, and Reactions




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3: Stoichiometry: Mass, Formulas, and Reactions Page 82

Search

Atoms




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↑ Previous: Questions and Problems

3

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3: Stoichiometry: Mass, Formulas, and Reactions

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Previous: Questions and Problems

3

Stoichiometry

Mass, Formulas, and Reactions

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Chemistry

1: Particles of Matter: Measurement and the Tools of Science

0 of 6 results in this section

ANCIENT UNIVERSE The colors of the more than 10,000 galaxies in this image give us a glimpse into the universe as it existed about 13 billion years ago. This image was taken by NASA's Hubble Space Telescope.

PARTICULATE REVIEW

Atoms and Molecules: What's the Difference?

In Chapter 1 we explore how chemists classify different kinds of matter, from elements to compounds to mixtures. Hydrogen and helium were the first two elements formed after the universe began. Chemists use distinctively colored spheres to distinguish **atoms** of different elements in their drawings and models. For example, hydrogen is almost always depicted as white.

- How many of the following particles are shown in this image?
 - Hydrogen **atoms**?
 - Hydrogen molecules?
 - Helium **atoms**?
- Are molecules composed of **atoms**, or are **atoms** composed of molecules?

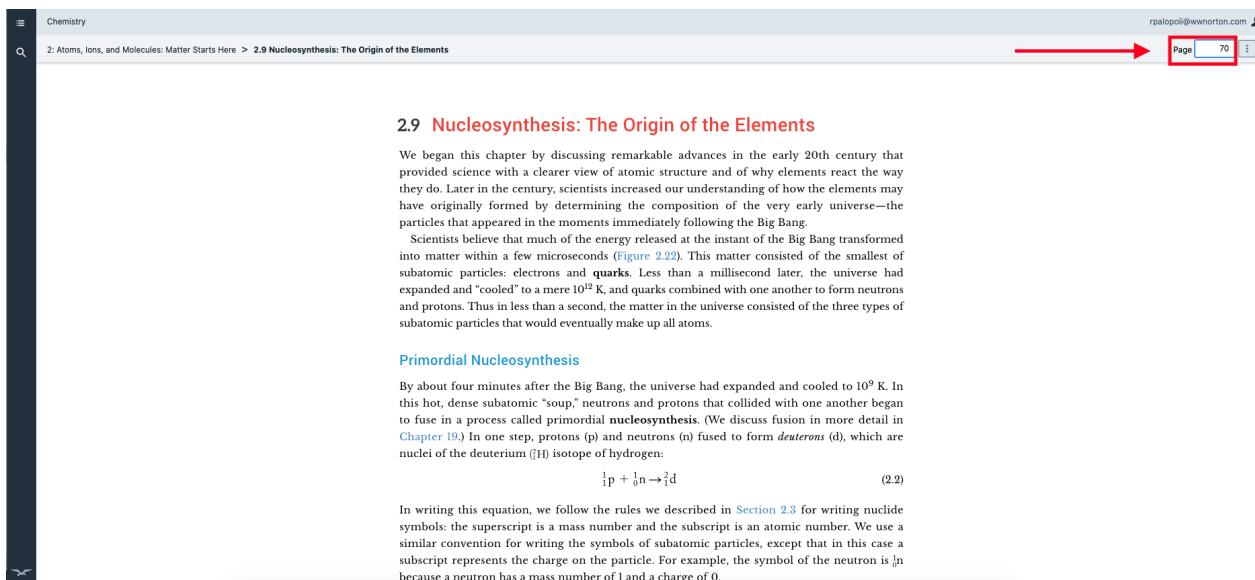
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Learning Outcomes

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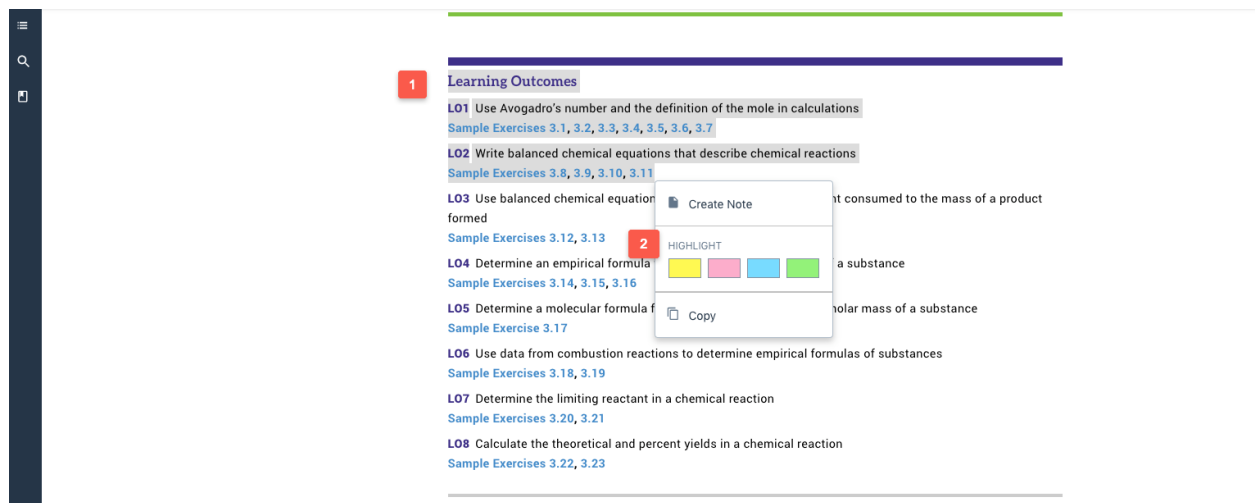


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Click **Delete Highlight**

6: Properties of Gases: The Air We Breathe

Notebook

- Classify the products as elements, compounds, or a mixture.
(Review Sections 1.1, 1.2, and 3.3 if you need help.)

SHOW ANSWER

1 Learning Outcomes

LO1 Distinguish gases from liquids and solids

LO2 Measure pressure and convert it to standard units to quantify it

Sample Exercises 6.1, 6.2

LO3 Calculate changes in the volume of a gas using the ideal gas law, the combined gas law, and the ideal gas law

Sample Exercises 6.3, 6.4, 6.5, 6.6

LO4 Use balanced chemical equations to determine the amount of a product by using the stoichiometric coefficients and the ideal gas law

Sample Exercises 6.8, 6.9

LO5 Calculate the density and molar mass of a gas

Sample Exercises 6.10, 6.11

LO6 Determine the mole fraction of a gas in a mixture

Sample Exercises 6.12, 6.13, 6.14

LO7 Use kinetic molecular theory to explain the behavior of gases

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6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O_2 to metabolize nutrients.

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1. Unlike the volume occupied by a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.
2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.
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4. Gases are typically much less dense than liquids or solids. One indicator of this large difference is that gas densities are expressed in grams per liter but liquid densities are expressed in grams per milliliter. The density of dry air at 20°C at typical atmospheric pressure is 1.20 g/L, for example, whereas the density of liquid water under the same conditions is 1.00 g/mL—more than 800 times greater than the density of dry air.

These four observations about gases are consistent with the idea that the particles of a gas (be they molecules or atoms) are further apart than the particles in solids and liquids. The larger

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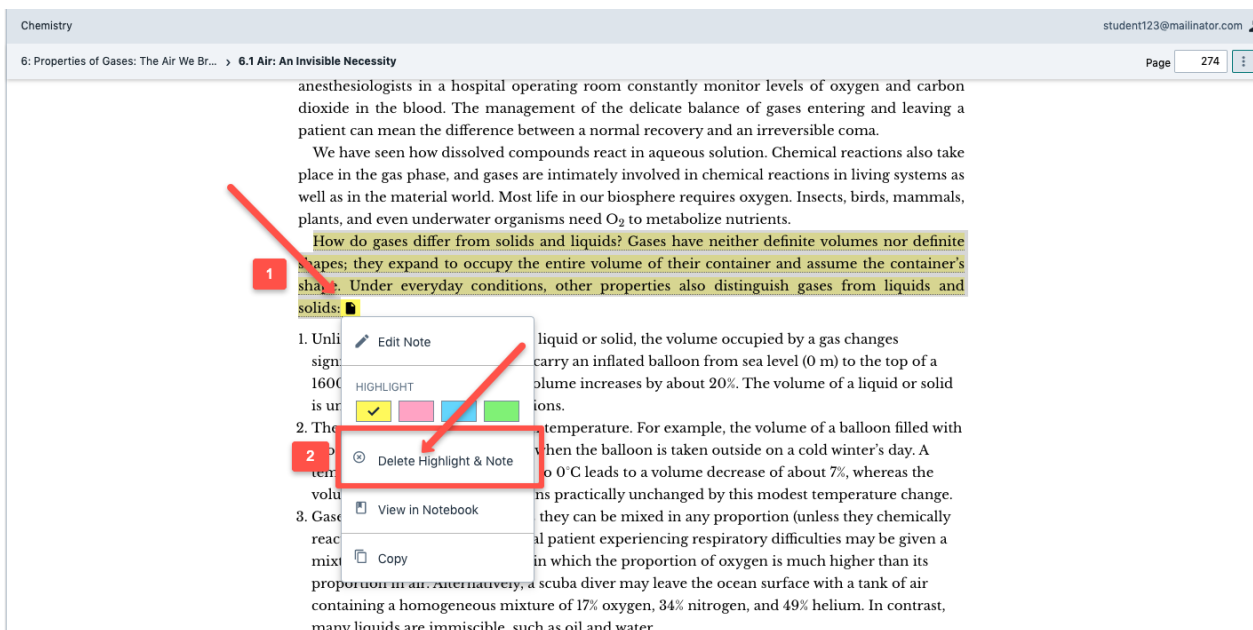
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- 2. The density of a gas is much lower than that of a liquid or solid. The density of a gas is also much more sensitive to temperature changes than that of a liquid or solid.
- 3. Gases can be mixed in any proportion (unless they chemically react with each other). A patient experiencing respiratory difficulties may be given a mixture of gases in which the proportion of oxygen is much higher than its proportion in air. Alternatively, a scuba diver may leave the ocean surface with a tank of air containing a homogeneous mixture of 17% oxygen, 34% nitrogen, and 49% helium. In contrast, many liquids are immiscible, such as oil and water.



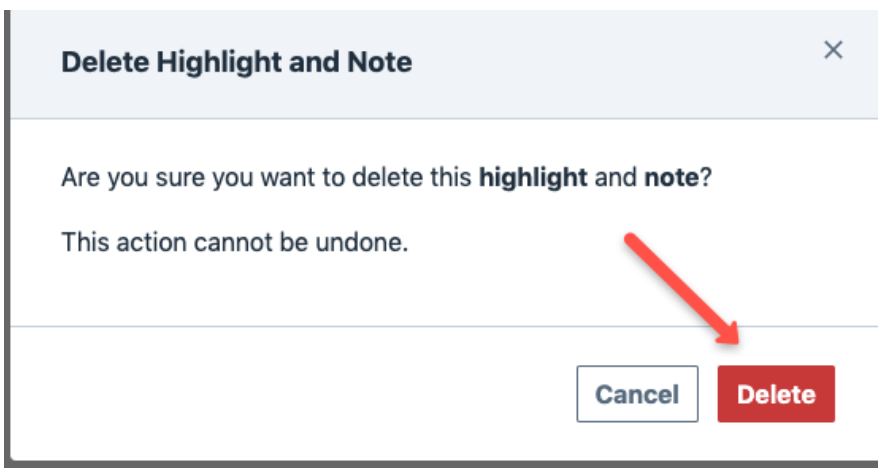
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1: Particles of Matter: Measurement and the Tools of Science

Page 2

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Measurement and the Tools of Science

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1: Particles of Matter: Measurement and the Tools of Science

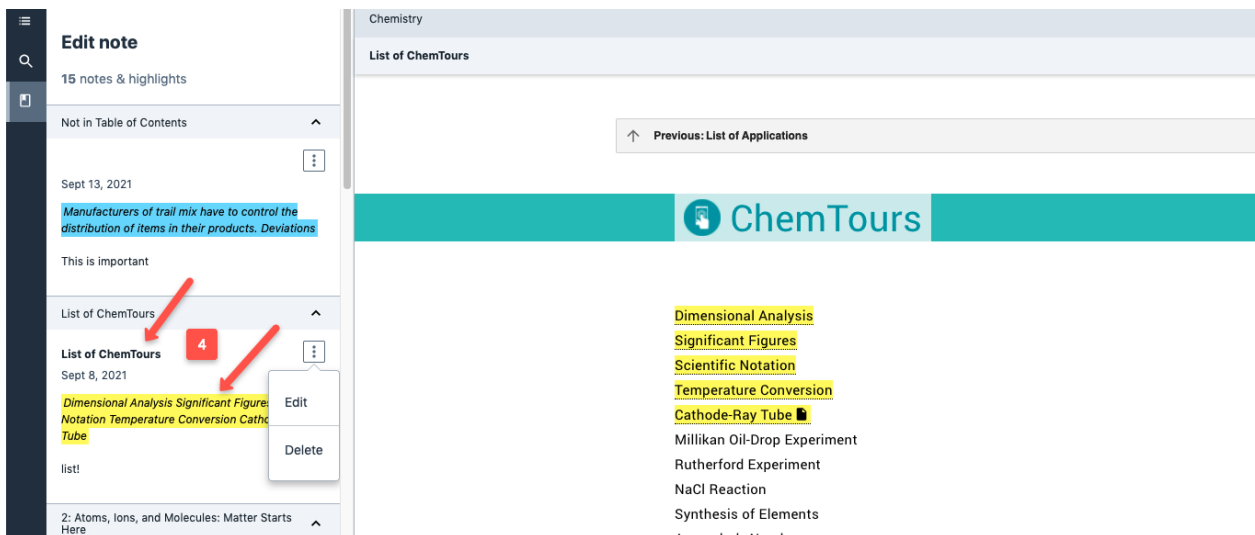
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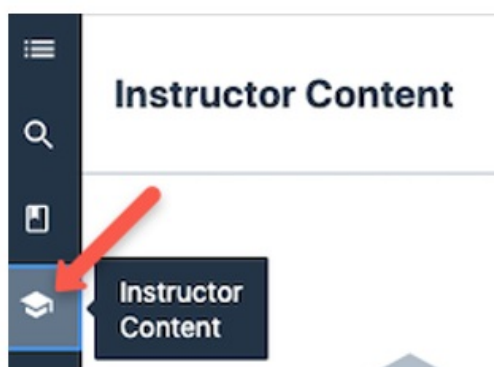
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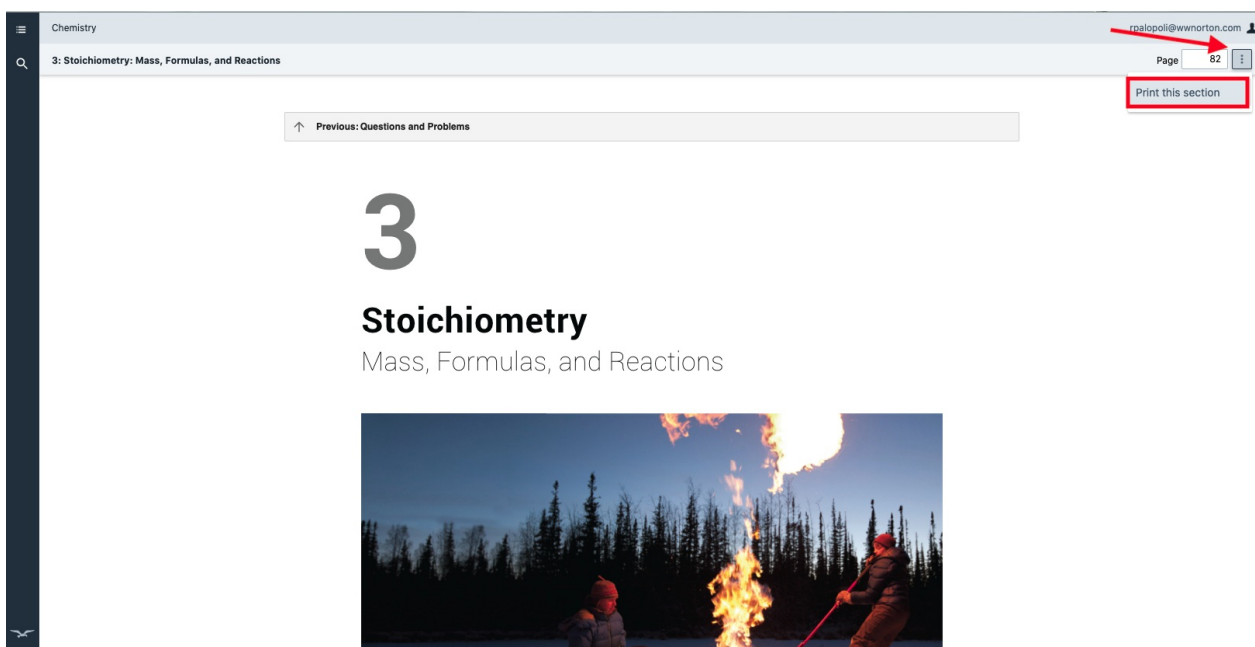
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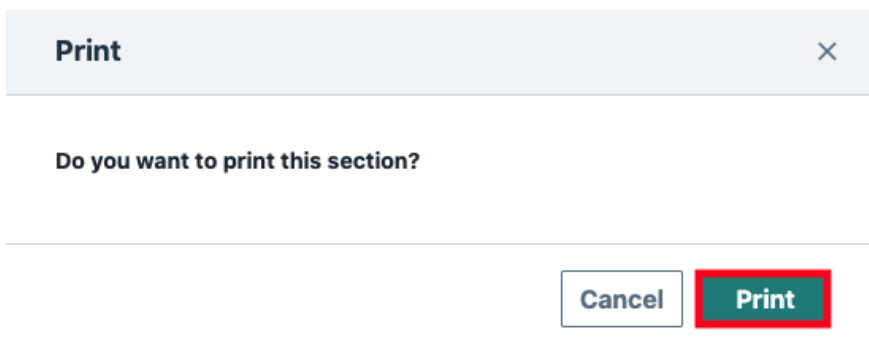
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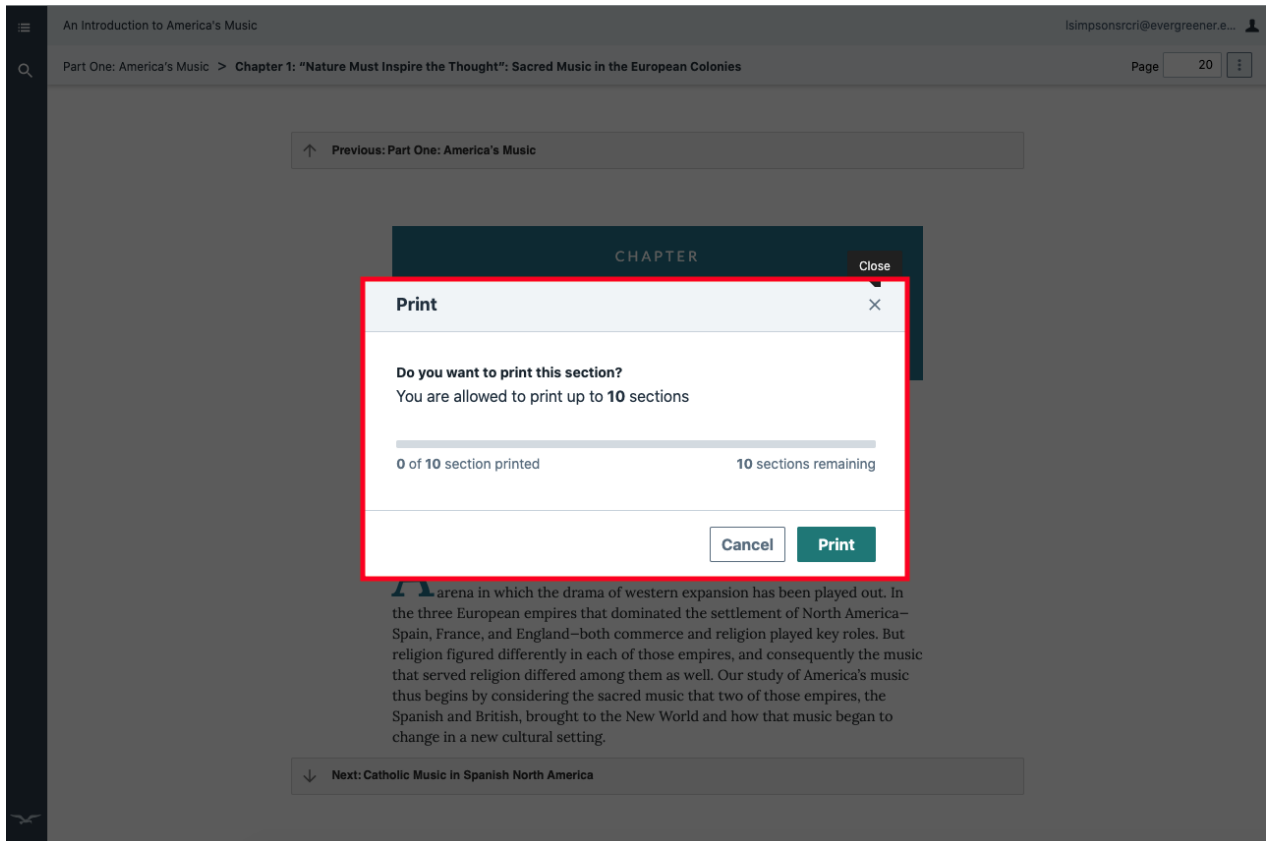


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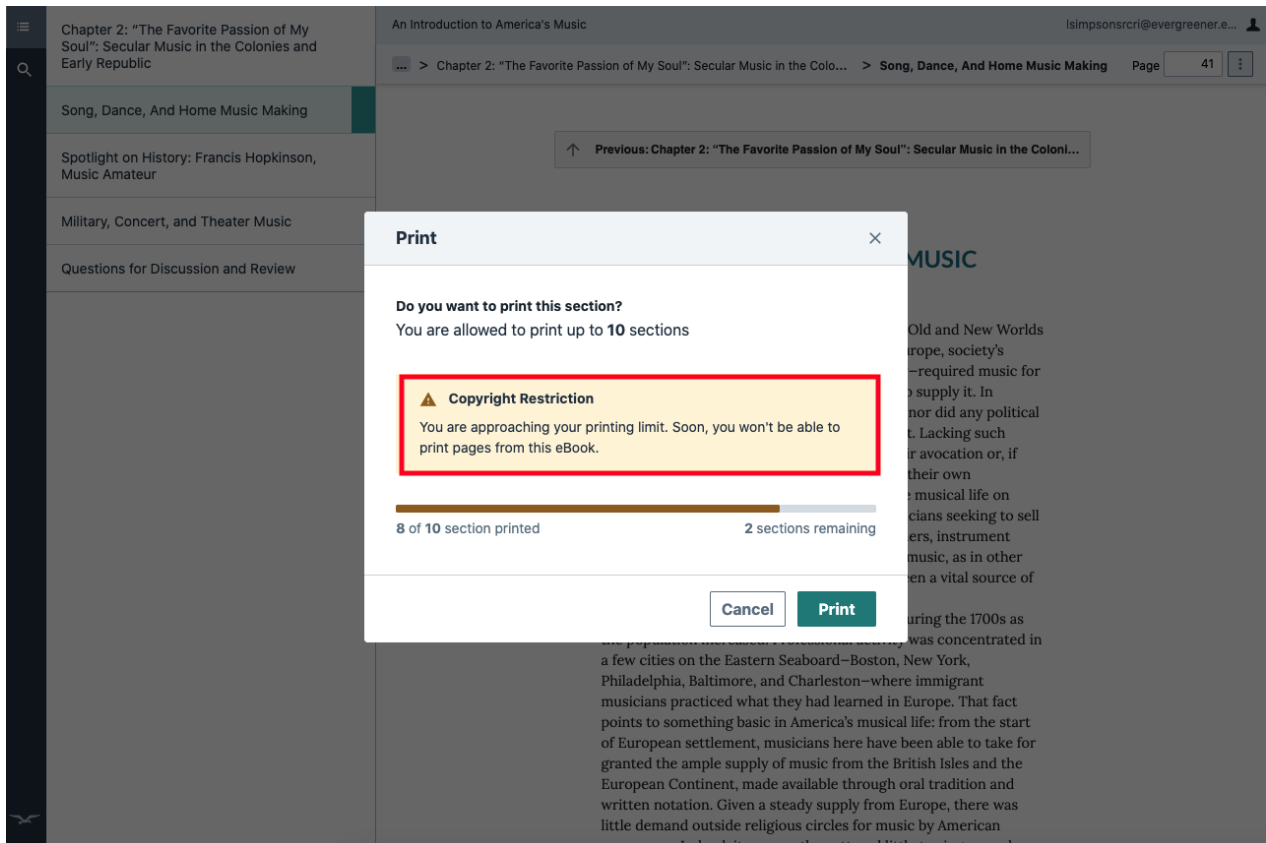
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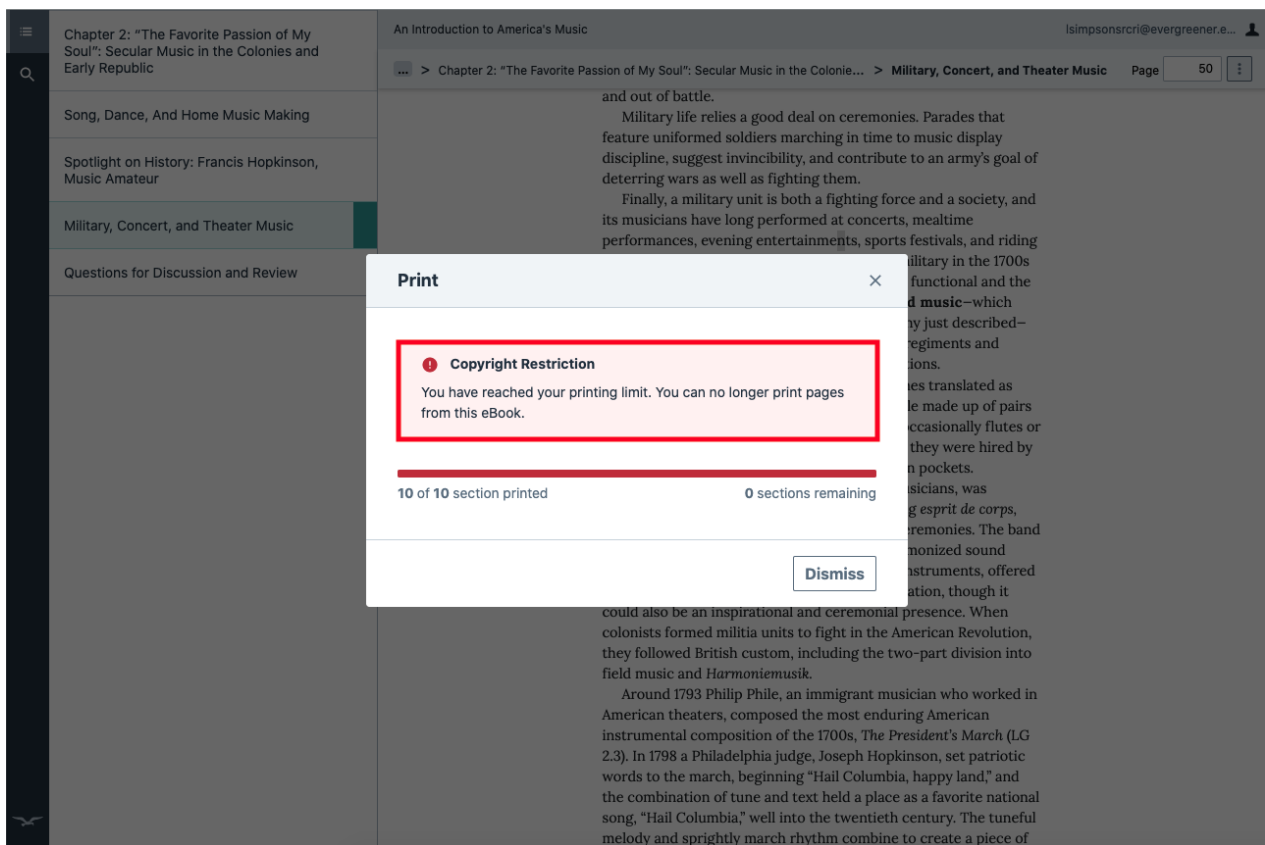
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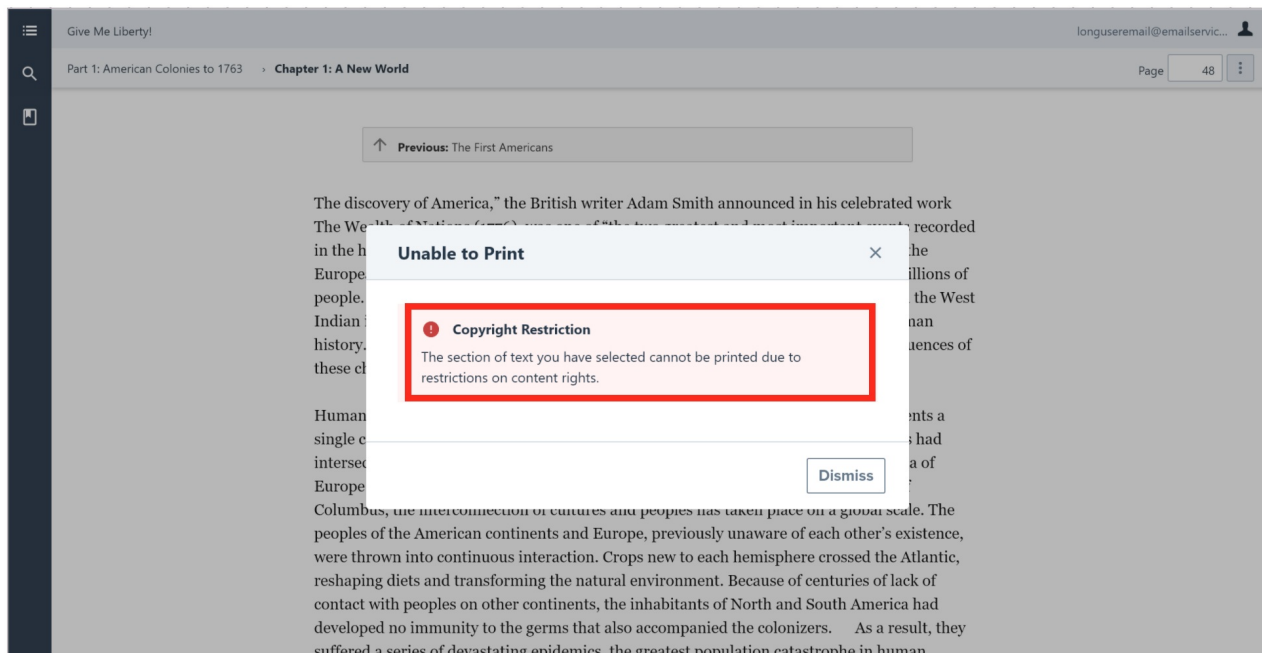


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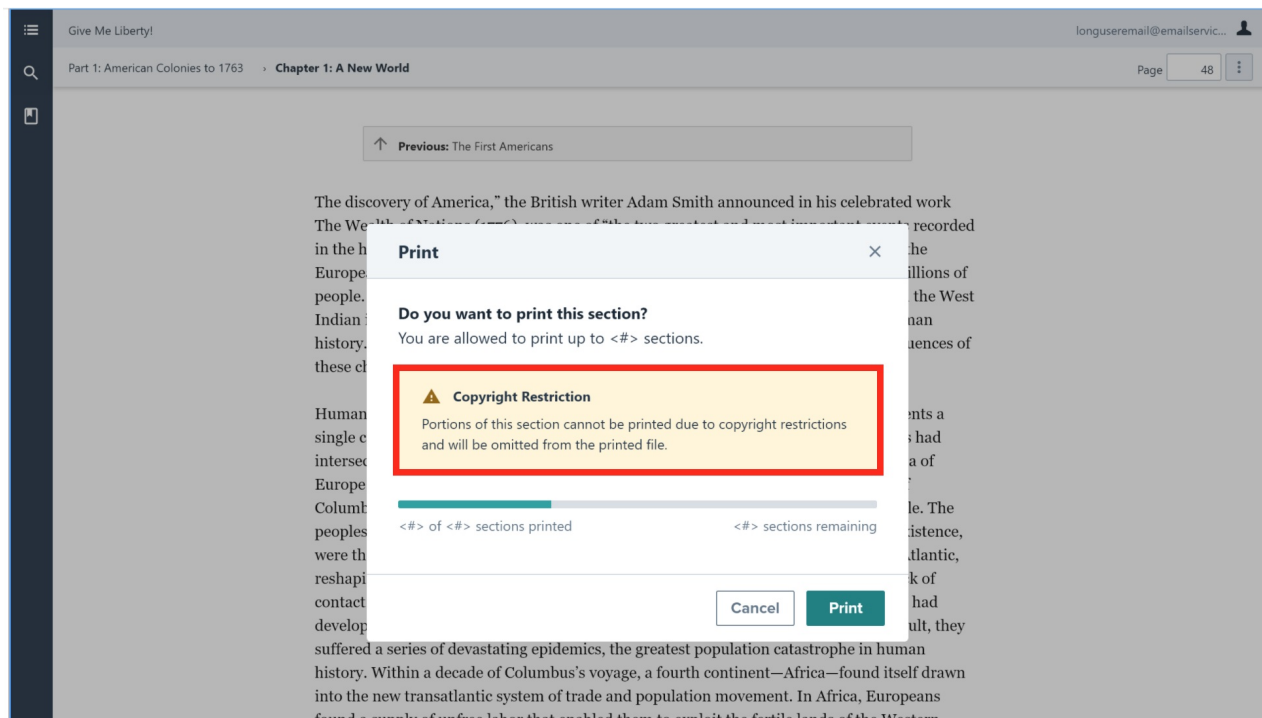
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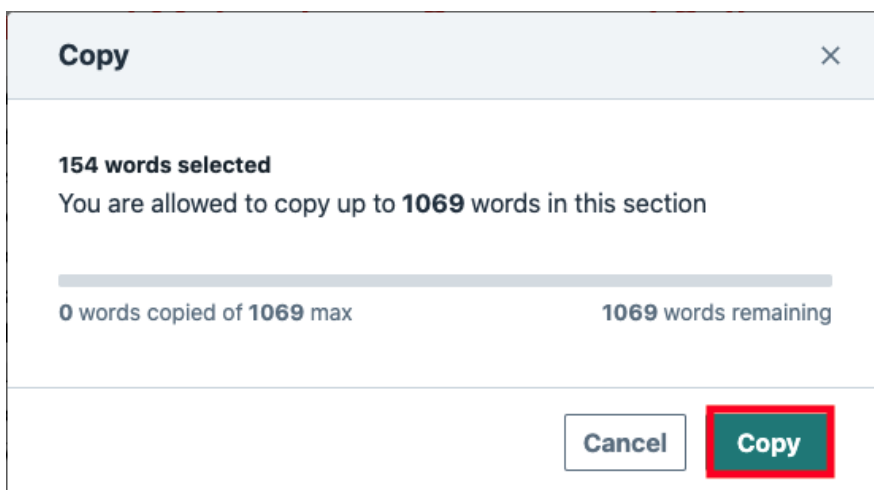
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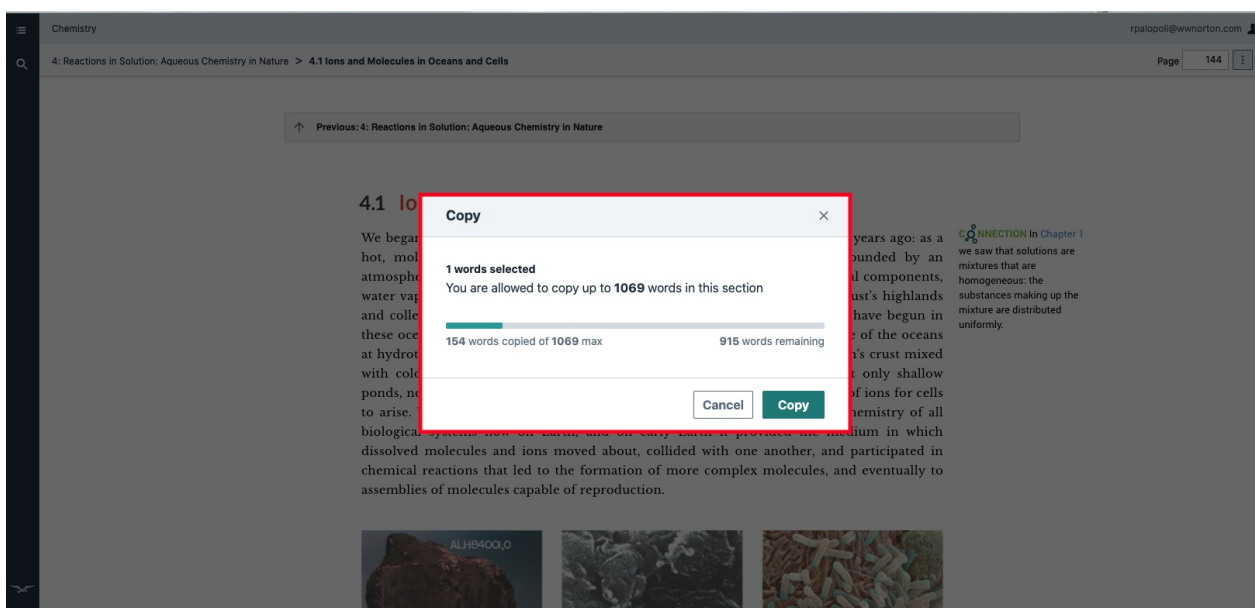
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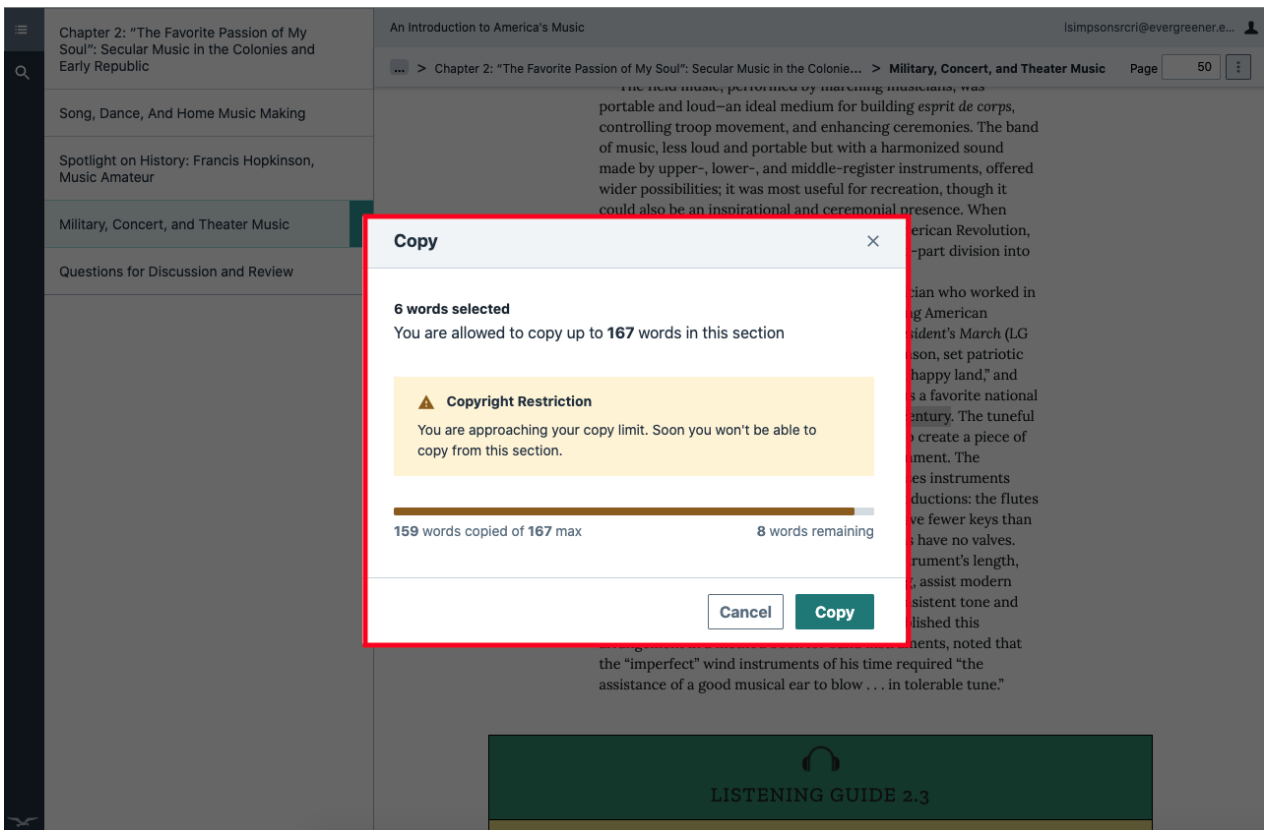
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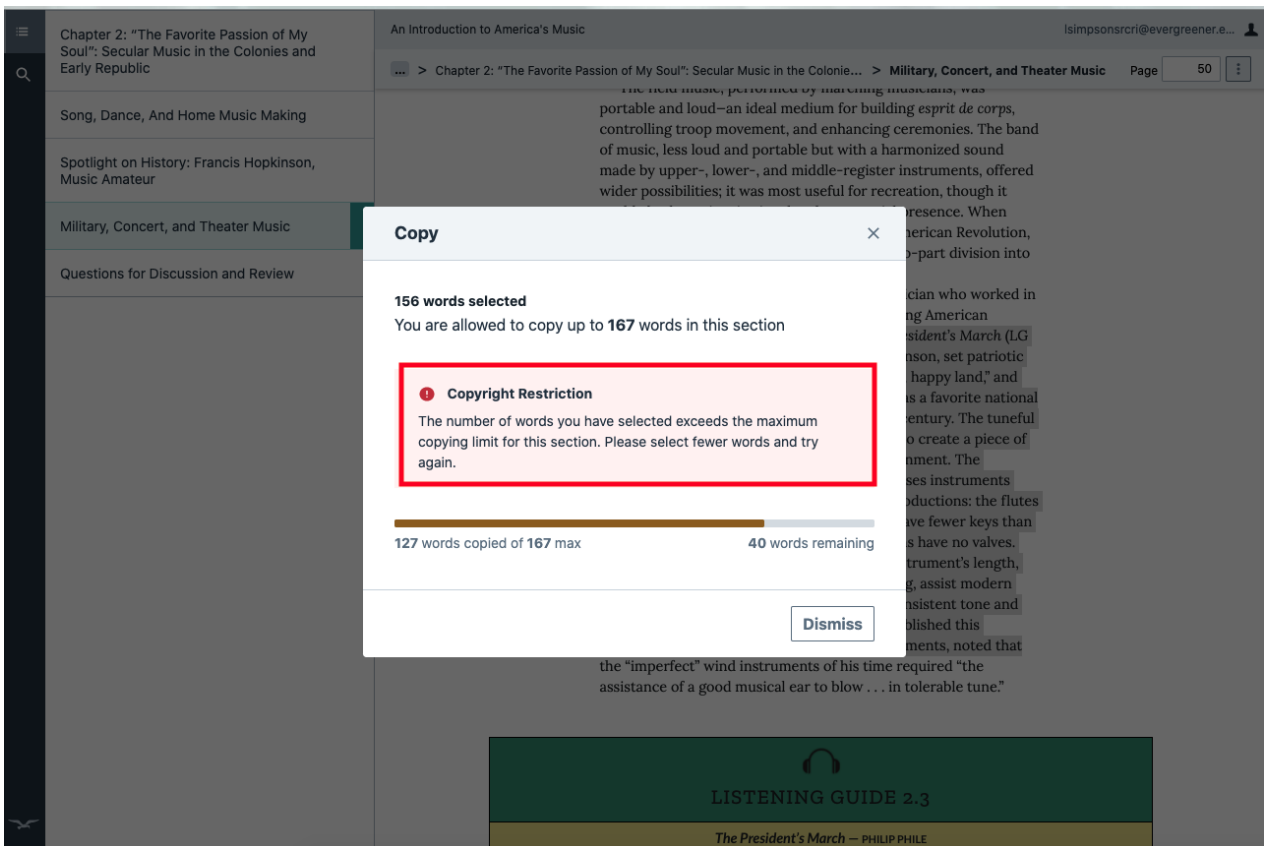
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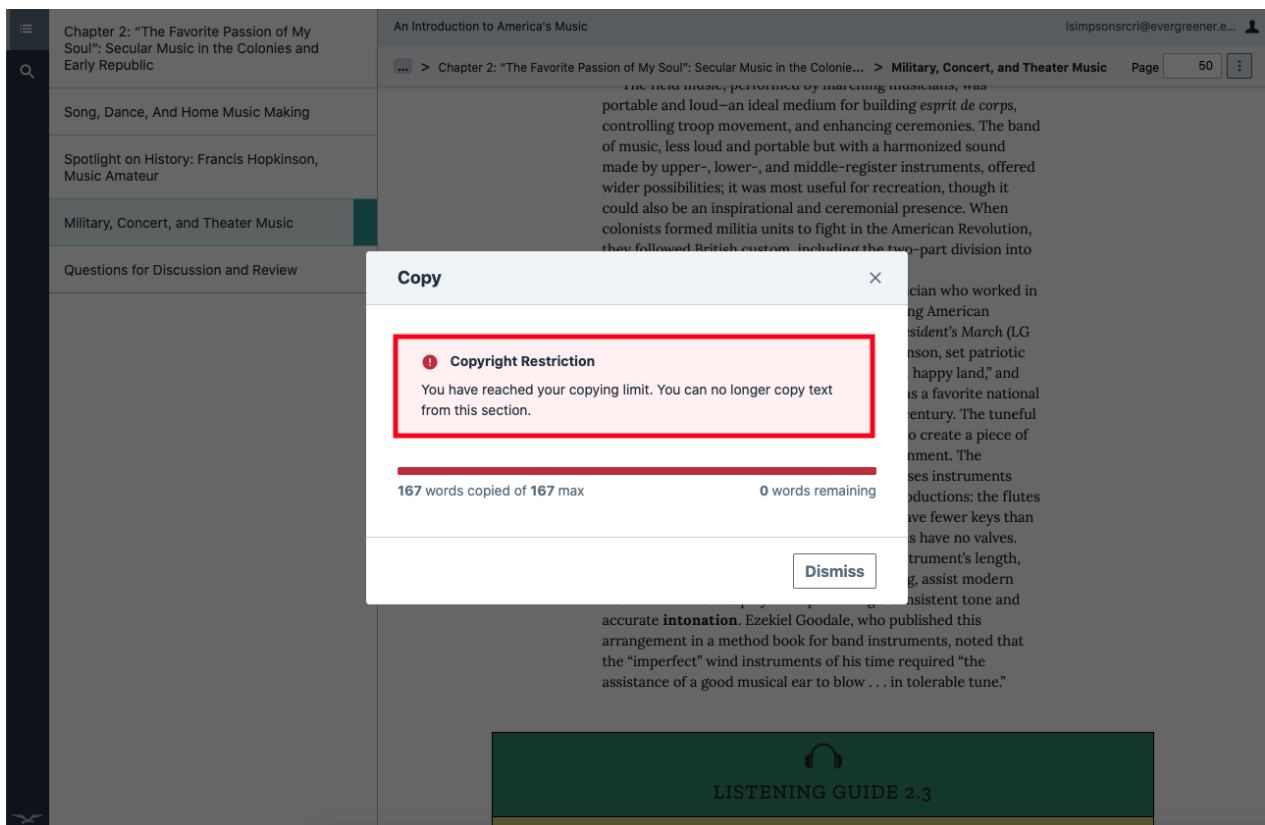
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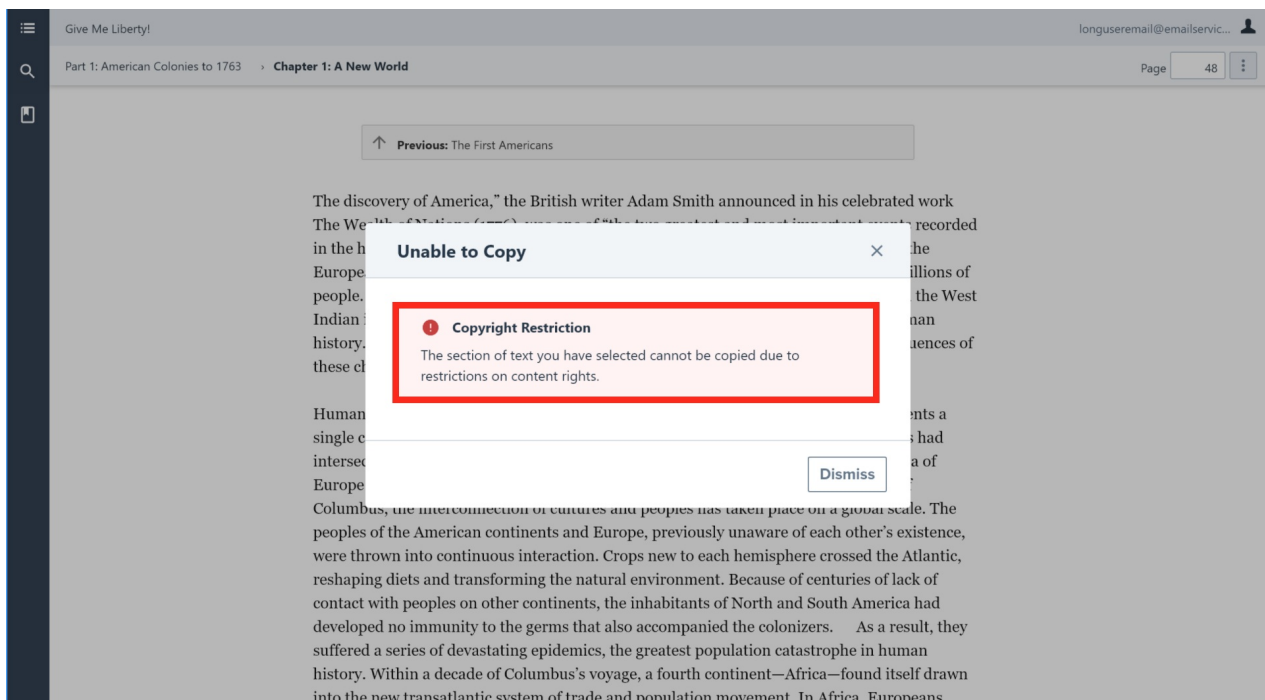


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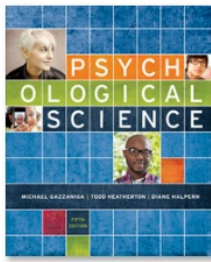


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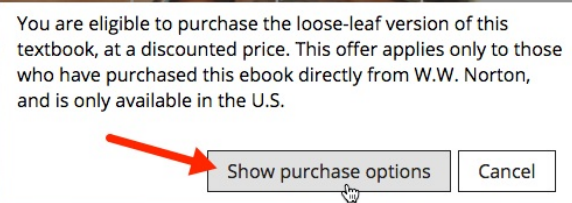


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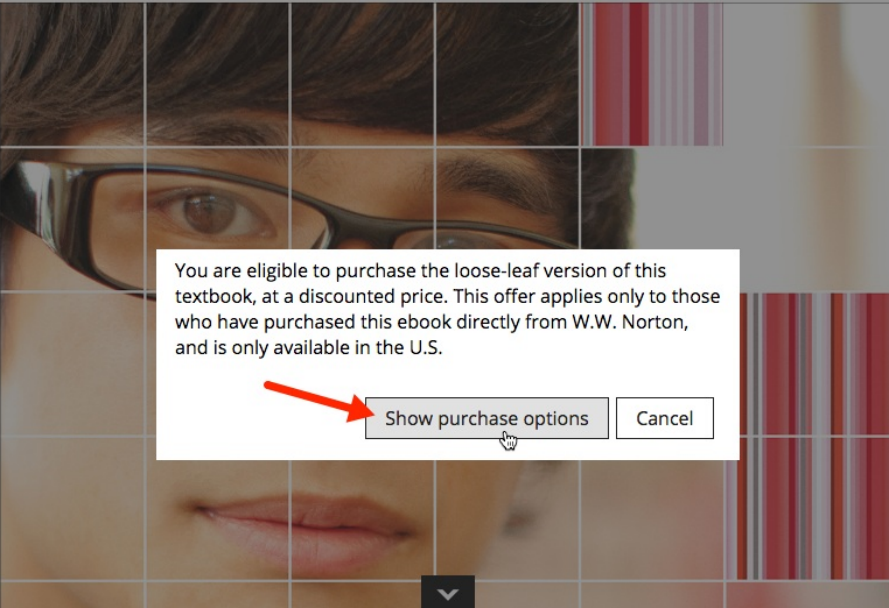
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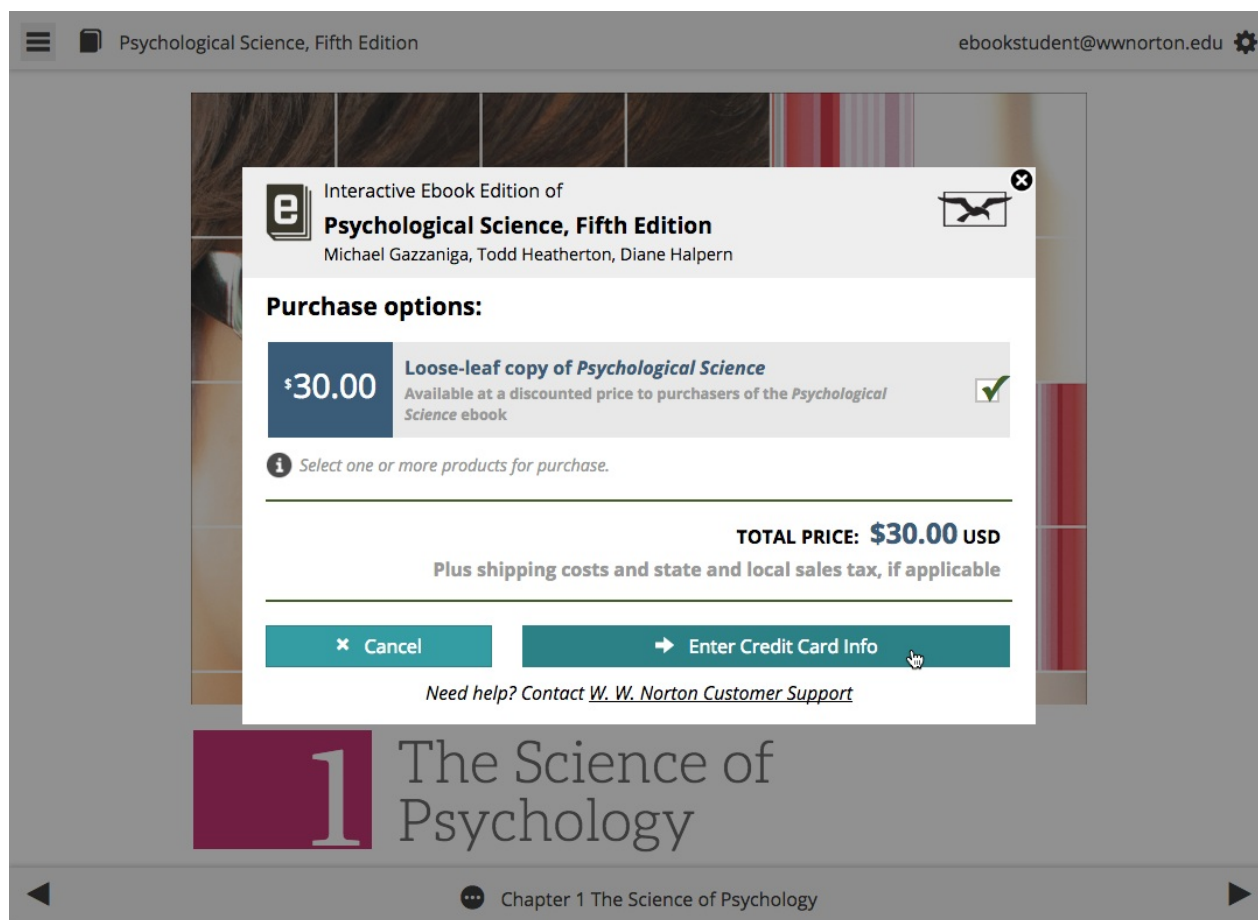
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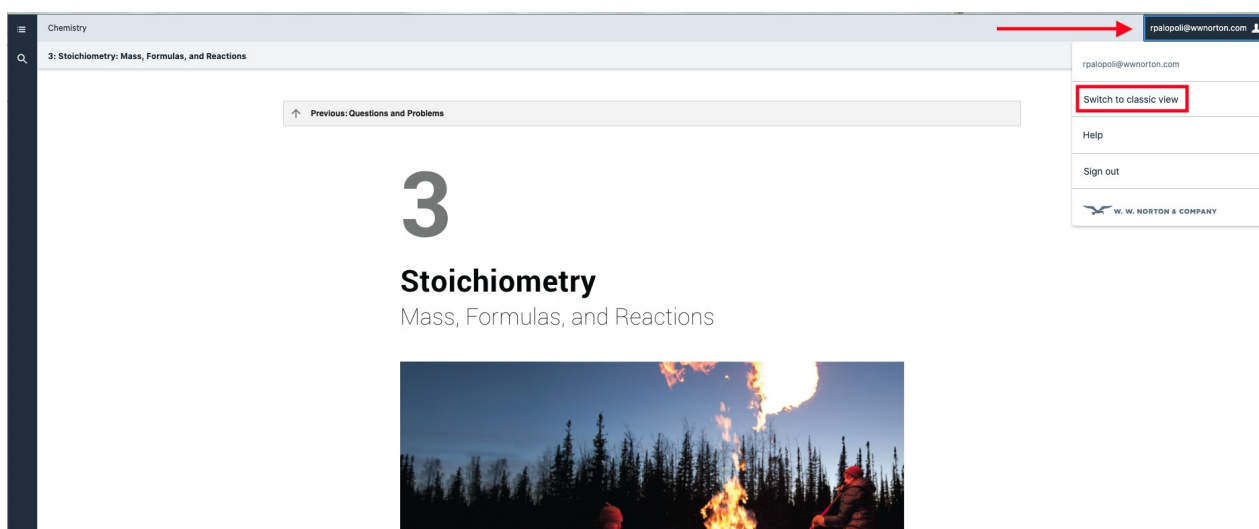
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