

Shared Instructor and Student Functionality

Last Modified on 12/01/2025 5:07 pm EST

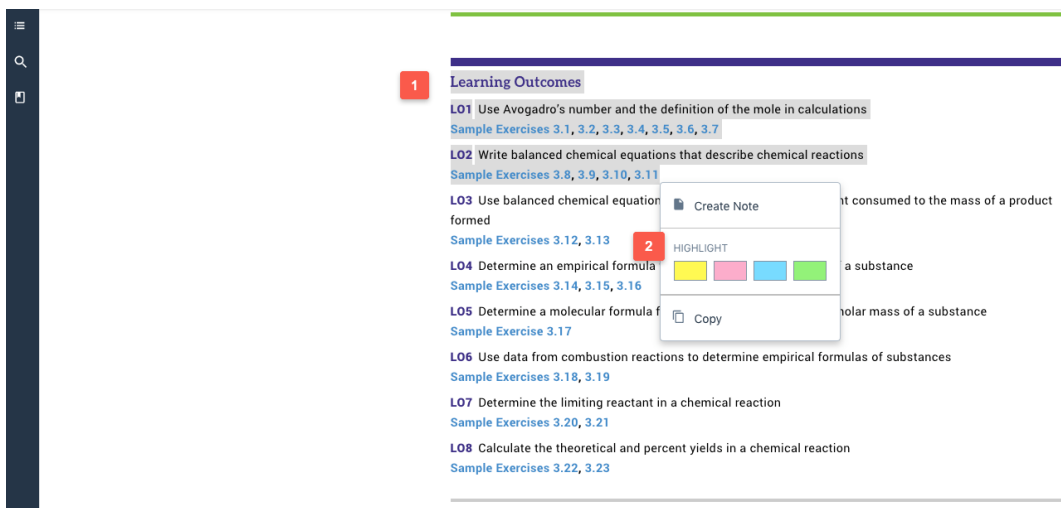
The Norton Ebook Reader has functionality that is common to both instructors and students. This page details the highlighting, annotation, bookmarking, audio narration, and printing and copying capabilities that are available to both instructors and students.

Hide All Answers

How do I highlight text?

Personal Highlights

To highlight text, use the cursor to select the text you would like to highlight and the **Context Menu** will appear. Select the color in which you would like the text highlighted: yellow, pink, blue, or green.



Removing Highlights

To remove a highlight, use the cursor to select the text from which you would like to remove the highlighting and the **Context Menu** will appear.

Click **Delete Highlight**

The screenshot shows a digital textbook page titled "6: Properties of Gases: The Air We Breathe". A sidebar on the left contains a "Notebook" icon. The main text area contains a list of learning outcomes (LO1 to LO7) and sample exercises. A red box labeled "1" highlights the "Learning Outcomes" section. A context menu is open over the text, with a red box labeled "2" highlighting the "Delete Highlight" option. The menu also includes options for "Create Note", "HIGHLIGHT" (with color swatches), "View in Notebook", and "Copy". A "SHOW ANSWER" button is visible at the top right of the text area.

Select **Delete** and the highlighting will be removed from the selected text.

The dialog box is titled "Delete Highlight" and contains the text: "Are you sure you want to delete this **highlight**?" and "This action cannot be undone." At the bottom, there are two buttons: "Cancel" and "Delete". A red arrow points to the "Delete" button.

Please note: Instructor shared highlights are no longer available in the New Ebook Reader since the green highlighter is now accessible to all users in the new Ebook. The new [Instructor Content](#) functionality allows instructors to create, edit, and publish shareable content with students.

How do I create annotations?

Personal Annotations

To create an annotation that will only appear in your ebook, use the cursor to select the text you would like to annotate and the **Context Menu** will appear.

Click **Create Note**

Chemistry student123@mailinator.com

6: Properties of Gases: The Air We Breathe > 6.1 Air: An Invisible Necessity Page 274

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O_2 to metabolize nutrients.

1 How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids:

2

1. Unlike the volume occupied by a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.

2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.

3. Gases are **miscible**, which means they can be mixed in any proportion (unless they chemically

Create Note

HIGHLIGHT

Copy

Type your annotation into the text field and click the **Save** button save your annotation.

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HIGHLIGHT

NOTE

Important definition

Cancel Save

Click on the **Notebook** page icon to view notes in the Notebook

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6: Properties of Gases: The Air We Breathe > 6.1 Air: An Invisible Necessity Page 274

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4. Gases are typically much less dense than liquids or solids. One indicator of this large difference is that gas densities are expressed in grams per liter but liquid densities are expressed in grams per milliliter. The density of dry air at 20°C at typical atmospheric pressure is 1.20 g/L, for example, whereas the density of liquid water under the same conditions is 1.00 g/mL—more than 800 times greater than the density of dry air.

These four observations about gases are consistent with the idea that the particles of a gas (be they molecules or atoms) are further apart than the particles in solids and liquids. The larger

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List of ChemTours

2: Atoms, Ions, and Molecules: Matter Starts Here

4: Reactions in Solution: Aqueous Chemistry in Nature

6: Properties of Gases: The Air We Breathe

6: Properties of Gases: The Air We Breathe > 6.1 Air: An Invisible Necessity

Sept 22, 2021

Learning Outcomes LO1 Distinguish gases from liquids and solids LO2 Measure pressure and convert between the different units used to quantify it Sample Exercises 6.1, 6.2 LO3 Calculate changes in the volume, temperature, pressure, and number of moles of a gas by using the individual, combined, and ideal gas laws Sample Exercises 6.3, 6.4, 6.5, 6.6, 6.7 LO4 Use balanced chemical equations to relate the volume of a gas-phase reactant to the amount of a product by using the stoichiometry of the reaction and the ideal gas law Sample Exercises 6.8, 6.9 LO5 Calculate the density and molar mass of any gas Sample Exercises 6.10, 6.11 LO6 Determine the mole fraction and the partial pressure of a gas in a mixture Sample Exercises 6.12, 6.13, 6.14

6: Properties of Gases: The Air We Breathe > 6.1 Air: An Invisible Necessity

Sept 23, 2021

How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids:

Important definition

Answers to Selected End-of-Chapter Questions and Problems

How to Edit Annotations

Click the **notebook** page icon. The **Context Menu** will appear. Select **Edit Note**

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HIGHLIGHT

NOTE

This item will be on the exam!

Cancel Save

How to Delete Annotations

1. Select the **notebook page icon** on the annotation that you want to delete
2. Click **Delete Highlight & Note**

Chemistry

1: Particles of Matter: Measurement and the Tools of Science

Page 2

↑ Previous:

1

Particles of Matter

Measurement and the Tools of Science

Chemistry

1: Particles of Matter: Measurement and the Tools of Science

↑ Previous:

1

Particles of Matter

Measurement and the Tools

1. This is the **total number** of notes and highlights
2. To **Edit** or **Delete** content select the 3 dots icon above the annotation or highlight
3. Annotations that you have created can be found under the highlights
4. Click on the **section title** to go directly to the page where an annotation or highlight is located.

Edit note

15 notes & highlights

Not in Table of Contents

Sept 13, 2021

Manufacturers of trail mix have to control the distribution of items in their products. Deviations

This is important

List of ChemTours

List of ChemTours

Sept 8, 2021

Dimensional Analysis Significant Figures Scientific Notation Temperature Conversion Cathode-Ray Tube

list!

2: Atoms, Ions, and Molecules: Matter Starts Here

Chemistry

List of ChemTours

Previous: List of Applications

ChemTours

Dimensional Analysis

Significant Figures

Scientific Notation

Temperature Conversion

Cathode-Ray Tube

Millikan Oil-Drop Experiment

Rutherford Experiment

NaCl Reaction

Synthesis of Elements

Avogadro's Number

Can I search my ebook for specific terms or page numbers?

Searching the Ebook

To search the text, select the magnifying glass from the left-hand side of the screen.

Chemistry

3: Stoichiometry: Mass, Formulas, and Reactions

Previous: Questions and Problems

3

Stoichiometry

Mass, Formulas, and Reactions

Page 82

Enter a term in the search field.

Chemistry

3: Stoichiometry: Mass, Formulas, and Reactions

Page 82

Search


Atoms

Type in the field above to search the book

3

Stoichiometry

Mass, Formulas, and Reactions



See the full book search results displayed below.

Chemistry

3: Stoichiometry: Mass, Formulas, and Reactions

Page 82

Search

Atoms

Cancel Search

Brief Contents

1: Particles of Matter: Measurement and the Tools of Science

Questions and Problems

"...heterogeneous. (Section 1.2) LO2 All matter consists of **atoms**, and we use chemical formulas consisting of atomic..."

2: Atoms, Ions, and Molecules: Matter Starts Here

Questions and Problems

"...of atomic structure. (Sections 2.1 and 2.2) LO2 **Atoms** consist of a nucleus containing protons and neutrons..."

3: Stoichiometry: Mass, Formulas, and Reactions

Questions and Problems

"...In a balanced chemical equation, the number of **atoms** of each element is the same on the reactant side..."

4: Reactions in Solution: Aqueous


0 results in this section

Previous: Questions and Problems

3

Stoichiometry

Mass, Formulas, and Reactions



Clicking on the search results will take you to that specific page in the ebook. Additionally, the keyword you entered will appear highlighted in the text, and you will see a note at the top of the page indicating how many times that word is used within the section.

Chemistry

1: Particles of Matter: Measurement and the Tools of Science

Page 2

Search

Atoms

Cancel Search

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"...In a balanced chemical equation, the number of **atoms** of each element is the same on the reactant side..."

4: Reactions in Solution: Aqueous

0 of 6 results in this section

ANCIENT UNIVERSE The colors of the more than 10,000 galaxies in this image give us a glimpse into the universe as it existed about 13 billion years ago. This image was taken by NASA's Hubble Space Telescope.

PARTICULATE REVIEW

Atoms and Molecules: What's the Difference?


In Chapter 1 we explore how chemists classify different kinds of matter, from elements to compounds to mixtures. Hydrogen and helium were the first two elements formed after the universe began. Chemists use distinctively colored spheres to distinguish **atoms** of different elements in their drawings and models. For example, hydrogen is almost always depicted as white.

How many of the following particles are shown in this image?

- Hydrogen **atoms**?
- Hydrogen molecules?
- Helium **atoms**?
- Are molecules composed of **atoms**, or are **atoms** composed of molecules?

SHOW ANSWER

Learning Outcomes

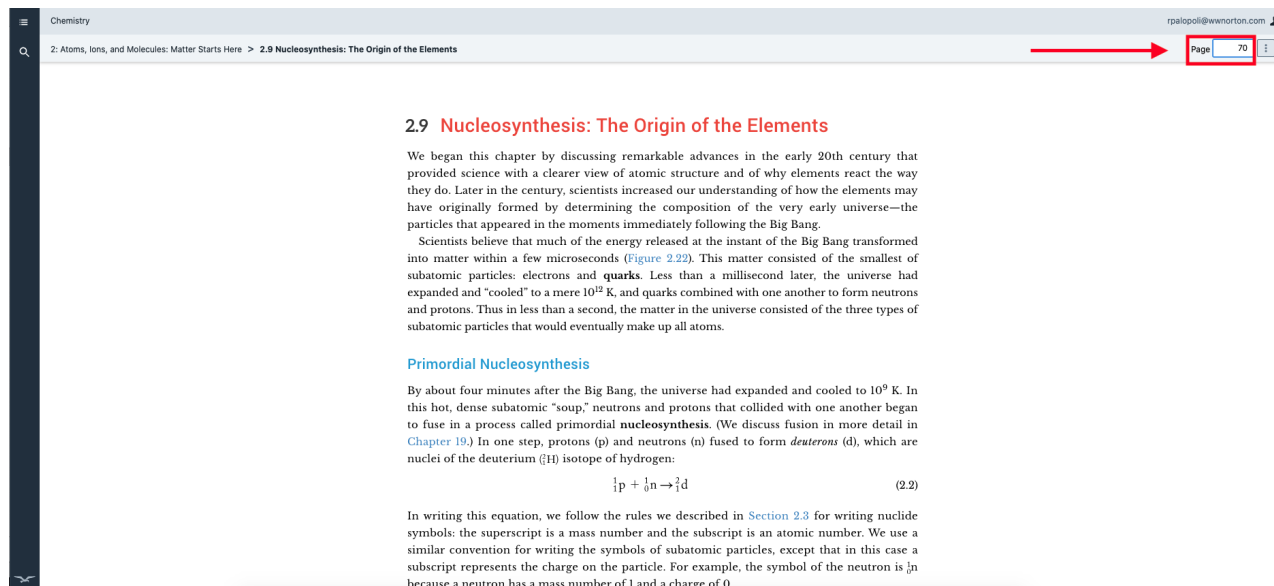


Please Note: These search results are for the entire book. Anytime the keyword you entered is displayed in the text, it

will show up here. If you would like to view the help notes on searching the Table of Contents, please click [here](#).

How do I search by page number?

You can search by a specific page number by inserting a number into the page field on the right-hand of the screen. This box will display the current page number you are viewing.



The screenshot shows the Norton Ebook Reader interface. At the top, the breadcrumb trail reads: "2: Atoms, Ions, and Molecules: Matter Starts Here > 2.9 Nucleosynthesis: The Origin of the Elements". On the right side of the header, there is a "Page" field with the number "70" entered. A red arrow points to this field. The main content area displays the title "2.9 Nucleosynthesis: The Origin of the Elements" in red. Below the title, the text describes the early 20th-century advances in atomic structure and the Big Bang. A section titled "Primordial Nucleosynthesis" follows, discussing the formation of matter from subatomic particles. A nuclear equation is shown:
$${}^1_1\text{p} + {}^1_0\text{n} \rightarrow {}^2_1\text{d} \quad (2.2)$$
 The text explains the notation for nuclides and subatomic particles.

After inserting a number in the page and select return on your keyboard, the ebook will take you to that page.

Can I read my ebook offline?

The newest version of the Norton Ebook Reader features the ability to read sections of your ebook offline.

Open any chapter of your ebook and select the **Table of Contents** icon.



The screenshot shows the Norton Ebook Reader interface. On the left side, there is a vertical sidebar with icons for "Table of Contents", "Search", "Notebook", and "Instructor Content". A red arrow points to the "Table of Contents" icon. The main content area displays the title "Chemistry: The Science in Context" and the chapter title "Chapter 1: Particles of Matter: Measurement and the Tools of Science". Below the chapter title, there is a "Previous" button labeled "Previous: Atomic Color Palette, Units, and Constants".

Click **Select Content for Offline Reading**

Chemistry: The Science in Context

Chapter 1: Particles of Matter: Measurement and the Tools of Science

↑ Previous: Atomic Color Palette, Units, and Constants

1

Particles of Matter

Measurement and the Tools of Science

Once offline reading is enabled, cached section **buttons** showing content available for offline reading will appear on the left as shown below.

Chemistry: The Science in Context

Chapter 9: Molecular Geometry: Shape Determines Function

↑ Previous: Summary

9

Molecular Geometry

Shape Determines Function

20.1 Molecular Structure and Functional Groups

20.2 Organic Molecules, Isomers, and Chirality

20.3 The Composition of Proteins

20.4 Protein Structure and Function

20.5 Carbohydrates

20.6 Lipids

20.7 Nucleotides and Nucleic Acids

20.8 From Biomolecules to Living Cells

Select the content you would like to make available for offline reading by selecting the **button** to the left of the chapter. You can also use the arrows to the right of the chapter title to view more detailed options when selecting content. Once you have finished selecting content, a progress bar will appear.

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CHEMISTRY

Chemistry: The Science in Context

Sixth Edition by Thomas Gilbert, Rein Kirss, Stacey Lowery Bretz, and Natalie Foster

Preparing for offline reading... (1/10 items)

Front Matter

Chapter 1: Particles of Matter: Measurement and the Tools of Science

Chapter 2: Atoms, Ions, and Molecules: Matter Starts Here

Chapter 3: Stoichiometry: Mass, Formulas, and Reactions

Chapter 4: Reactions in Solution: Aqueous Chemistry in Nature

Chapter 5: Properties of Gases: The Air We Breathe

Chemistry: The Science in Context

Chapter 1: Particles of Matter: Measurement a... > 1.1 How and Why

Previous: Chapter 1: Particles of Matter: Measurement and the Tools of Science

1.1 How and Why

For thousands of years, we humans have sought to better understand the world around us. For most of that time we resorted to mythological explanations of natural phenomena. Many once believed, for example, that the Sun rose in the east and set in the west because it was carried across the sky by a god driving a chariot propelled by winged horses.

In recent times we have been able to move beyond such fanciful accounts of natural phenomena to explanations based on observation and scientific reasoning. Unfortunately, this movement toward rational explanations has not always been smooth. Consider, for example, the contributions of Galileo Galilei, the man Albert Einstein called the father of modern science. At the dawn of the 17th century, Galileo used advanced telescopes of his own design to observe the movement of the planets and their moons. He concluded that they, like Earth, revolved around the Sun. However, this view conflicted with a belief held by many religious leaders of his time that Earth was the center of the universe. In 1633

Please note: The more content you select, the longer it will take to make your selection available for offline reading.

When a section has been successfully cached, you will see the **Available Offline** tab as shown here and a check mark will appear next to the cached content.

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Select content for offline reading. Cancel

Front Matter

Chapter 1: Particles of Matter: Measurement and the Tools of Science

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Chemistry: The Science in Context

Chapter 1: Particles of Matter: Measurement a... > 1.1 How and Why

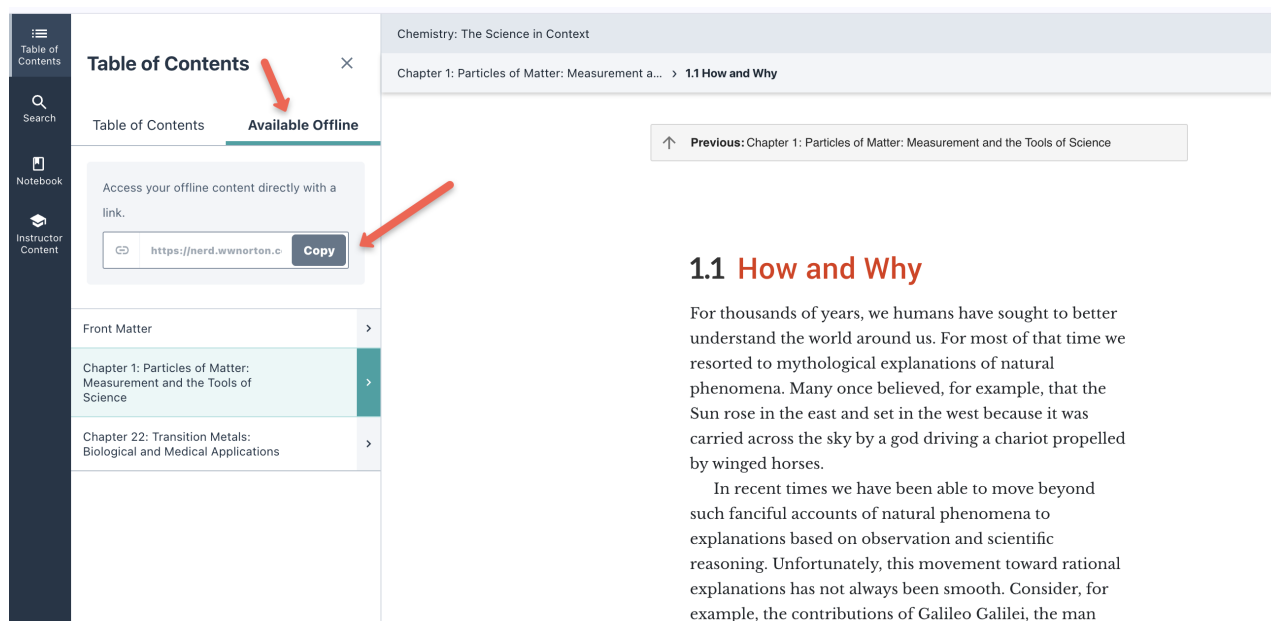
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Chemistry: The Science in Context

Chapter 1: Particles of Matter: Measurement a... > **1.1 How and Why**

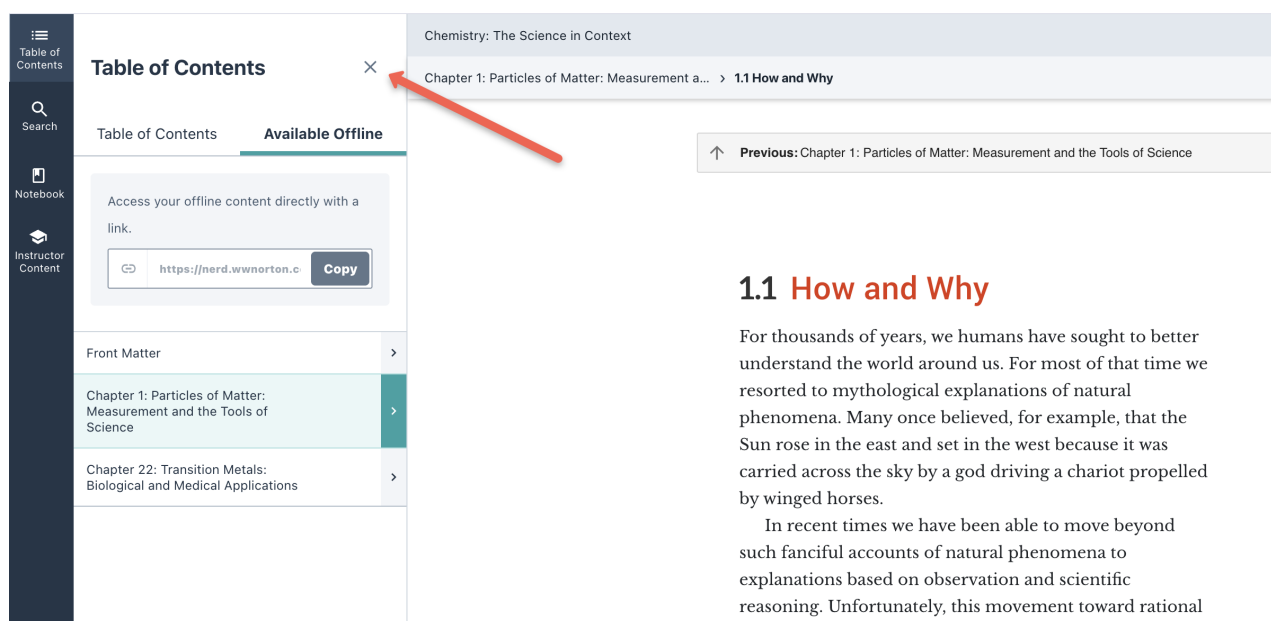
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Click on the X to return to the Table of Contents.



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Please Note: Media content (audio, video, animations, etc.) found in the ebook will only function with an active internet connection; those resources cannot be made available offline.

Highlights and Annotations

- In offline reading mode, you will only see notes and highlights for the content you've cached for offline reading.
- Notes and highlights cannot be created while in Offline Reading mode.

Parts of the ebook I previously made available offline are no longer available offline.

- Offline reading uses storage built into internet browsers. This means that when you make parts of your ebook available for offline reading, you aren't actually downloading or saving anything to your computer or mobile device.
- Sometimes, when your browsing history is cleared, the parts of the ebook you've made available offline will be cleared and will no longer be available for offline reading.
- To make these sections available for offline reading again, you'll need to get back online and repeat the same process you used to make those selections available offline the first time.

Does the ebook support audio narration?

Norton Ebooks and the Norton Ebook Reader are compatible both with screen readers and with browser extensions that enable text-to-speech functionality, such as the “Read Aloud” tool available on Google Chrome and Firefox. Apple and Google also offer text-to-speech solutions that you can use to read your ebook aloud on mobile devices.

Please click on one of the links below for more information:

- Apple iOS: [VoiceOver](#)
- Chrome: [Read Aloud: A Text to Speech Voice Reader](#)
- Google Android: [Google Text-to-speech](#)
- Firefox: [Read Aloud: A Text to Speech Voice Reader](#)

How do I print a specific section of the ebook?

To print a specific section, select the three dots next to the page number at the top.

The screenshot shows the Norton Ebook Reader interface. At the top, the page number '82' is displayed next to a three-dot menu icon. A red arrow points to this menu icon, and a red box highlights the 'Print this section' button that appears in the dropdown menu. The main content area displays the chapter title '3 Stoichiometry' and the subtitle 'Mass, Formulas, and Reactions'. Below the text is a photograph of two people in winter gear holding a large flaming torch in a snowy, wooded area at night.

Select 'print' from the confirmation box.

Print

×

Do you want to print this section?

Cancel

Print

Can I print the entire ebook?

The number of sections you are allowed to print varies by title.

When you select print on a section, you will see a popup window with a progress bar which indicates how many sections you have already printed and how many you have left to print.

An Introduction to America's Music

lsimpsonsrcrri@evergreener.e...

Part One: America's Music > Chapter 1: "Nature Must Inspire the Thought": Sacred Music in the European Colonies

Page 20

↑ Previous: Part One: America's Music

CHAPTER

Close

Print

×

Do you want to print this section?

You are allowed to print up to 10 sections

0 of 10 section printed

10 sections remaining

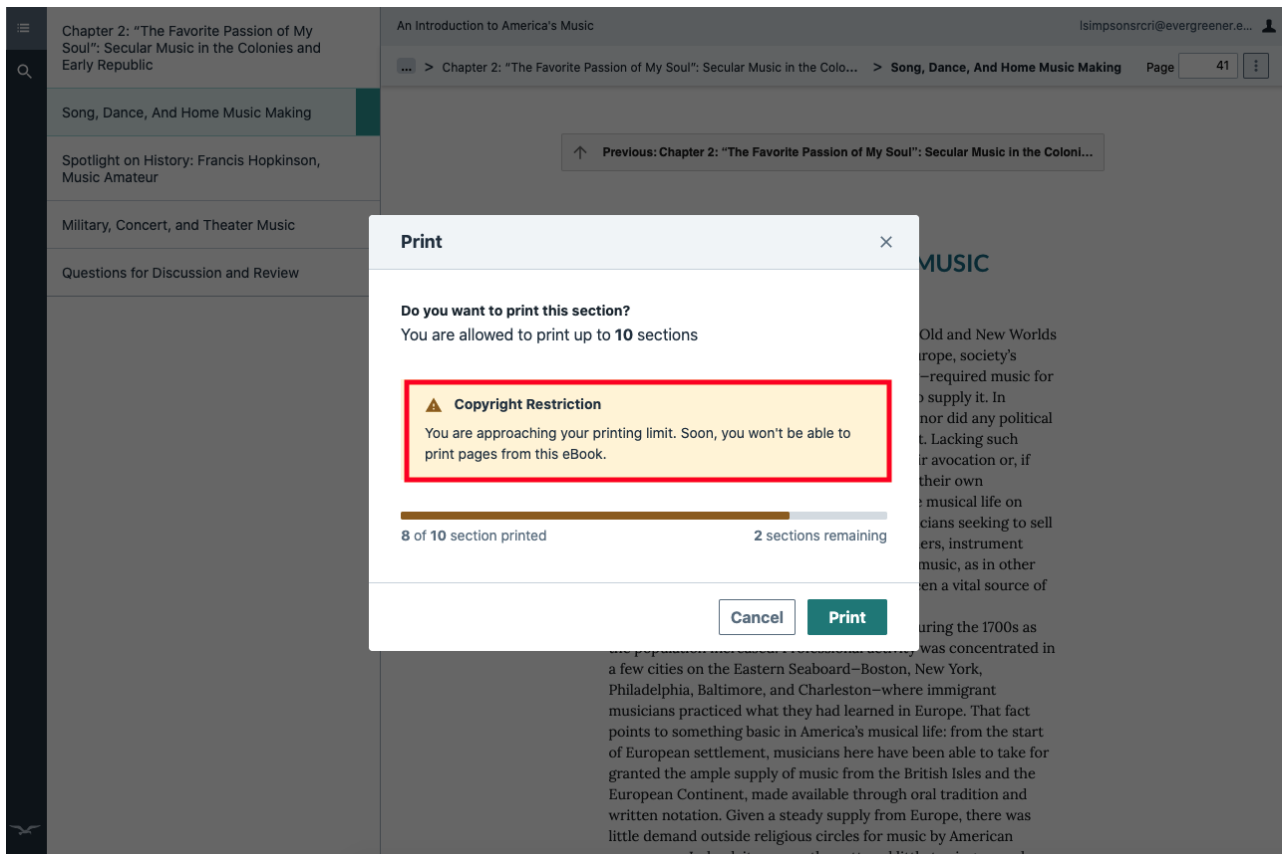
Cancel

Print

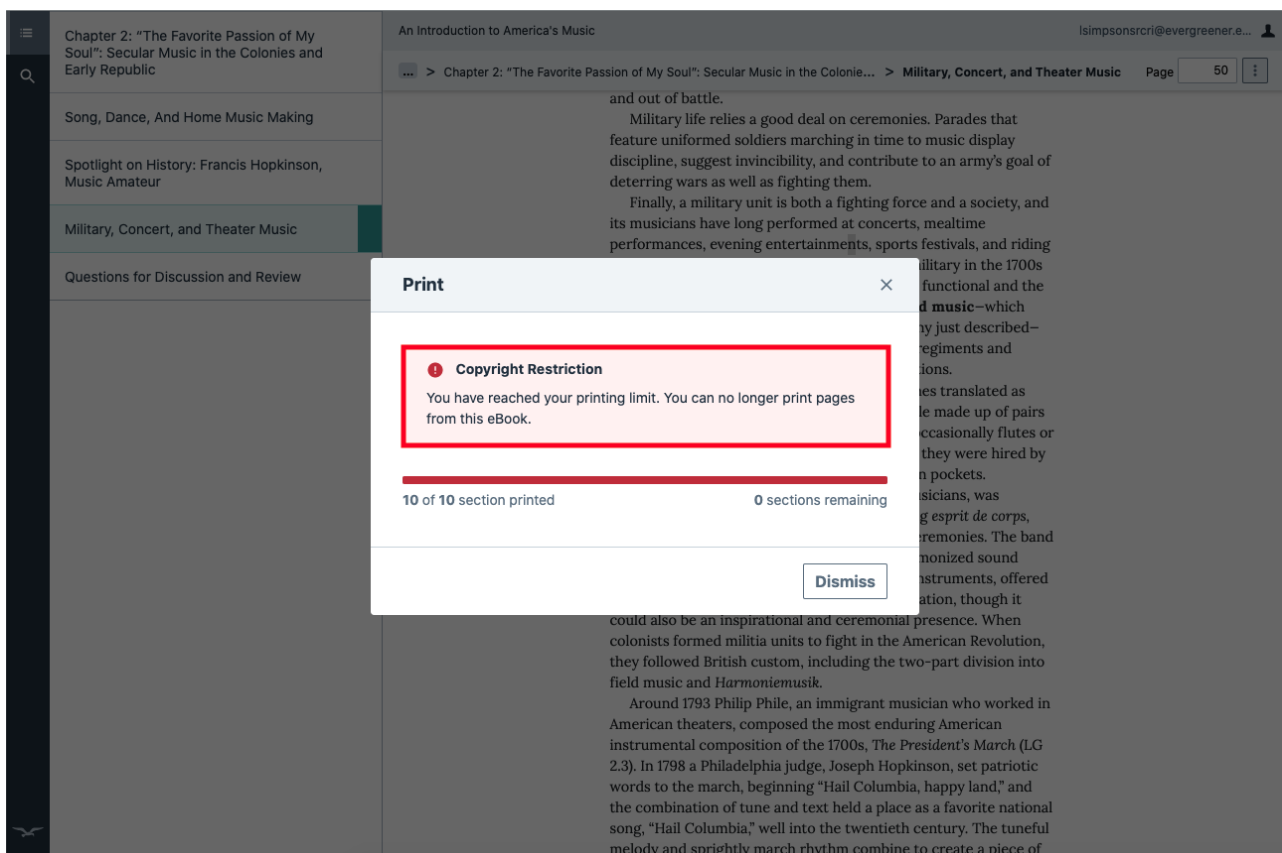
Larena in which the drama of western expansion has been played out. In the three European empires that dominated the settlement of North America—Spain, France, and England—both commerce and religion played key roles. But religion figured differently in each of those empires, and consequently the music that served religion differed among them as well. Our study of America's music thus begins by considering the sacred music that two of those empires, the Spanish and British, brought to the New World and how that music began to change in a new cultural setting.

↓ Next: Catholic Music in Spanish North America

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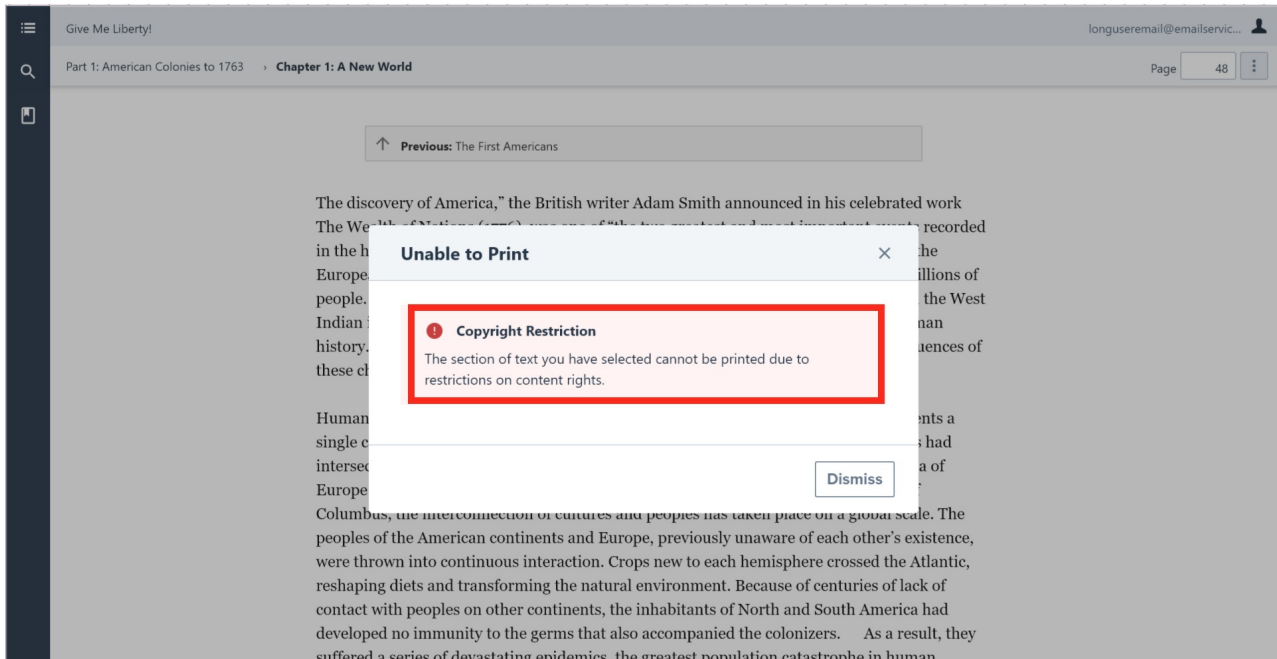
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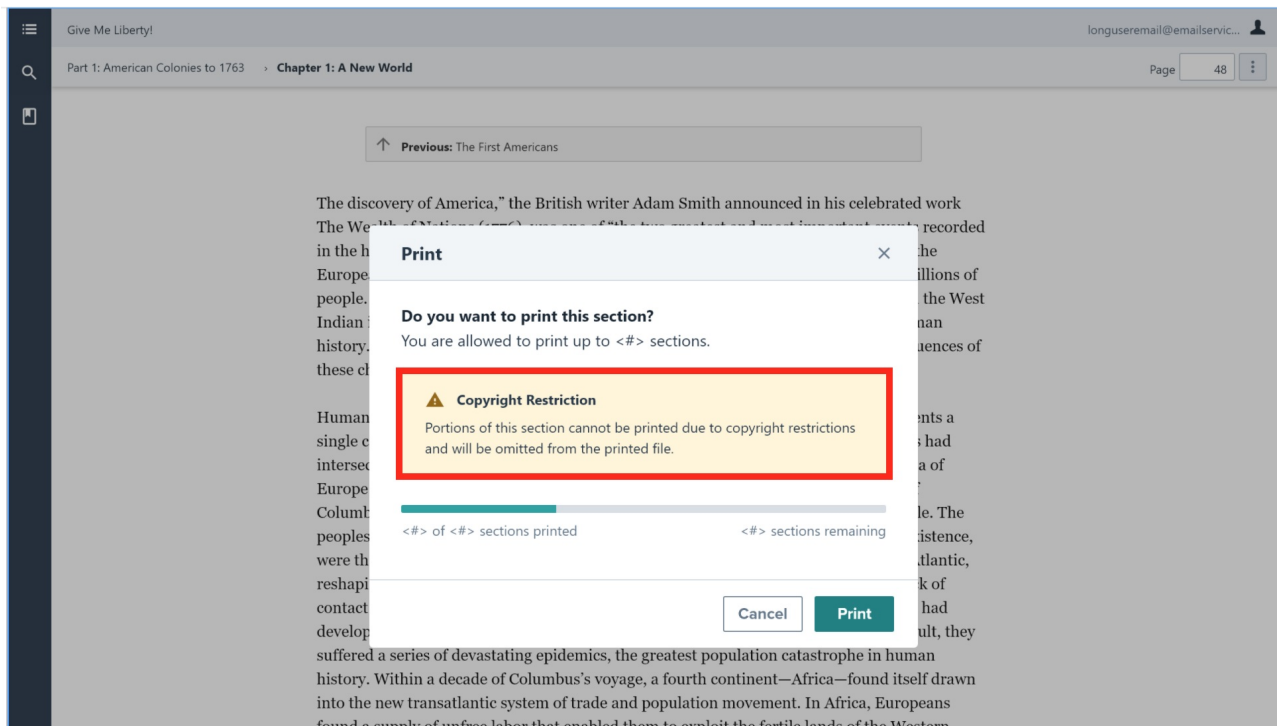
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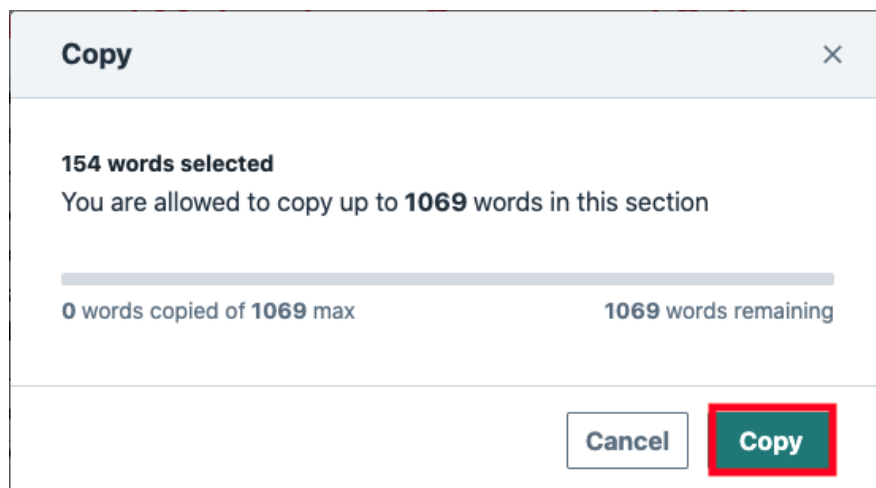
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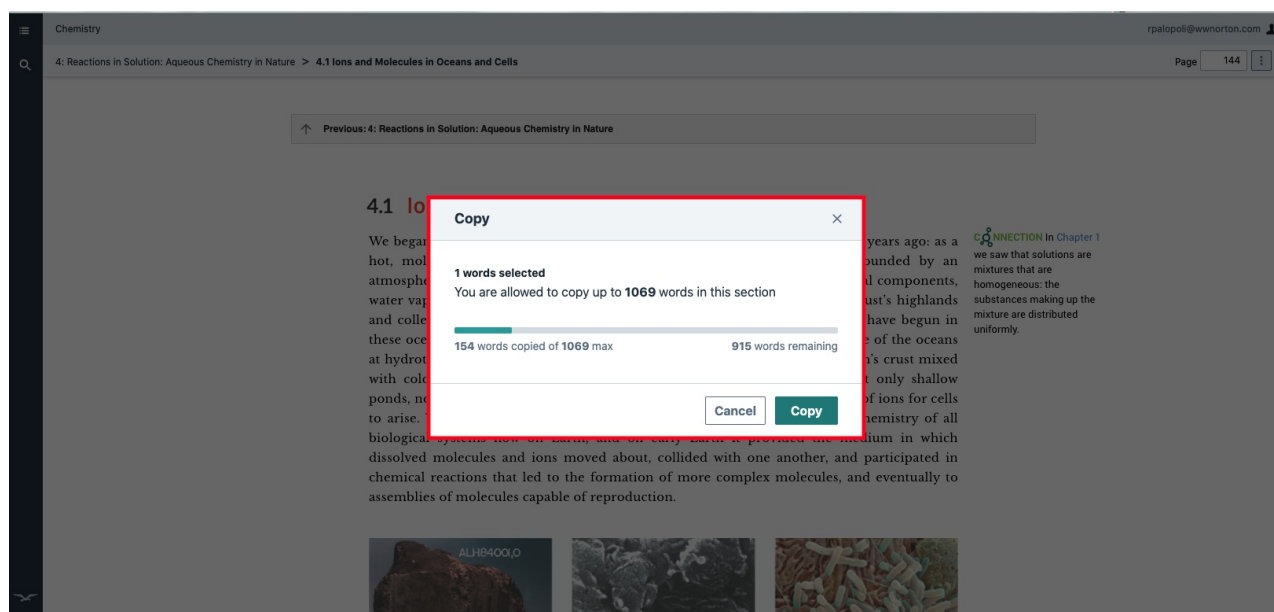
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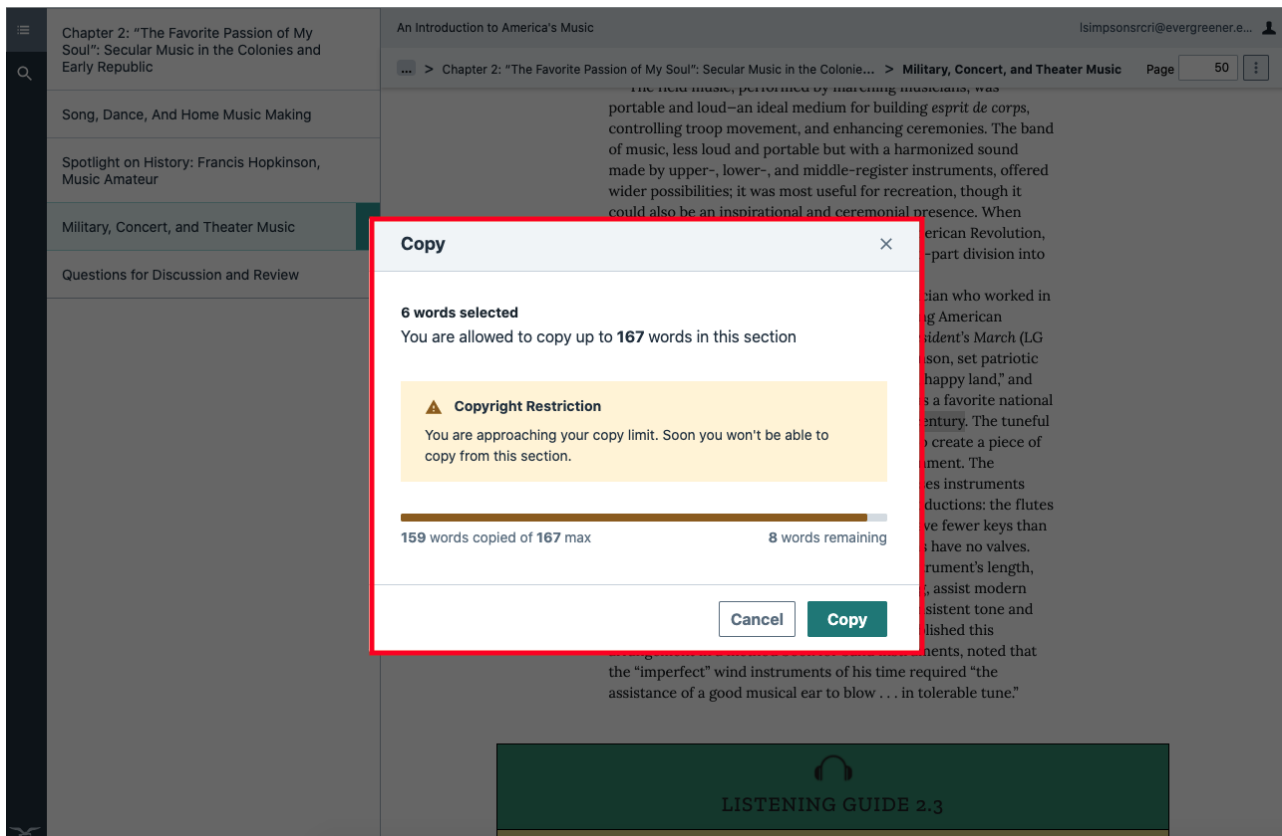
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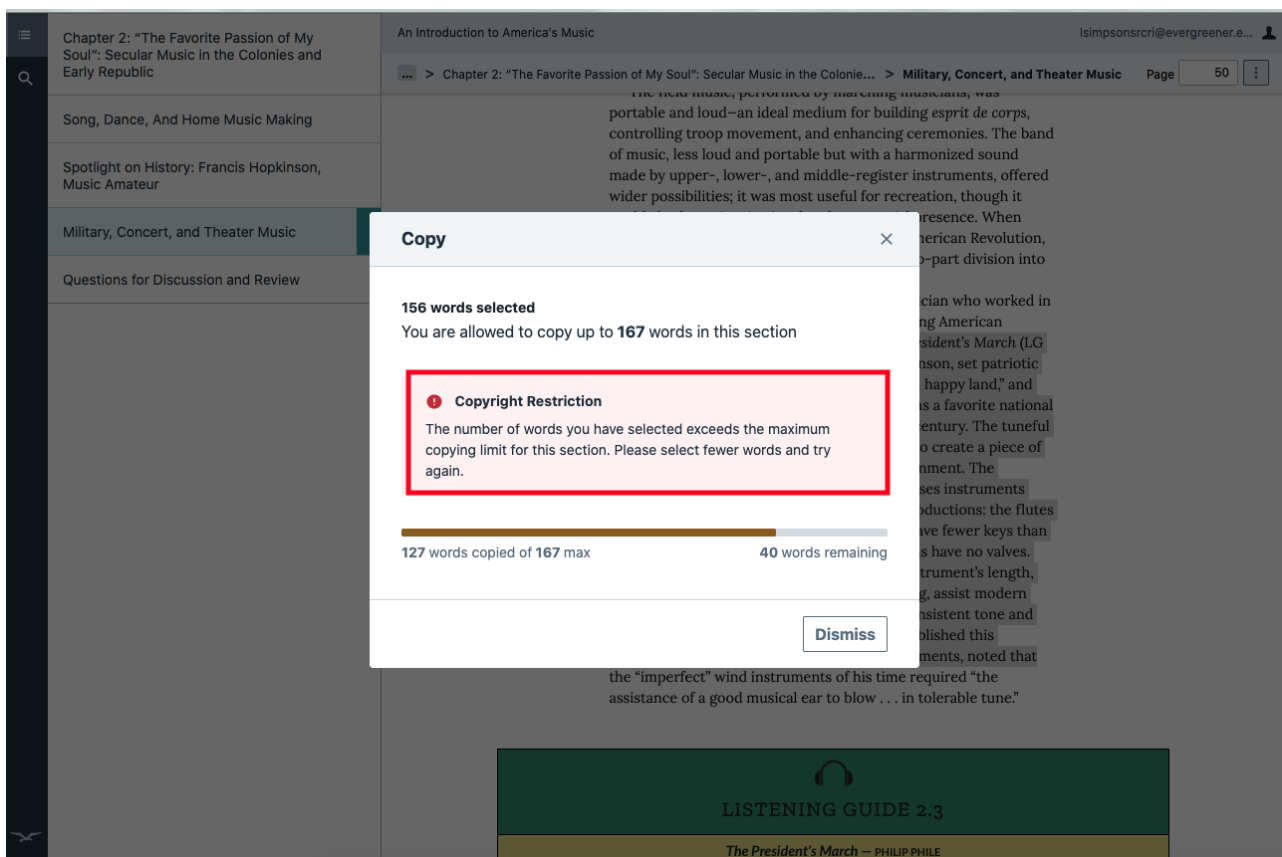
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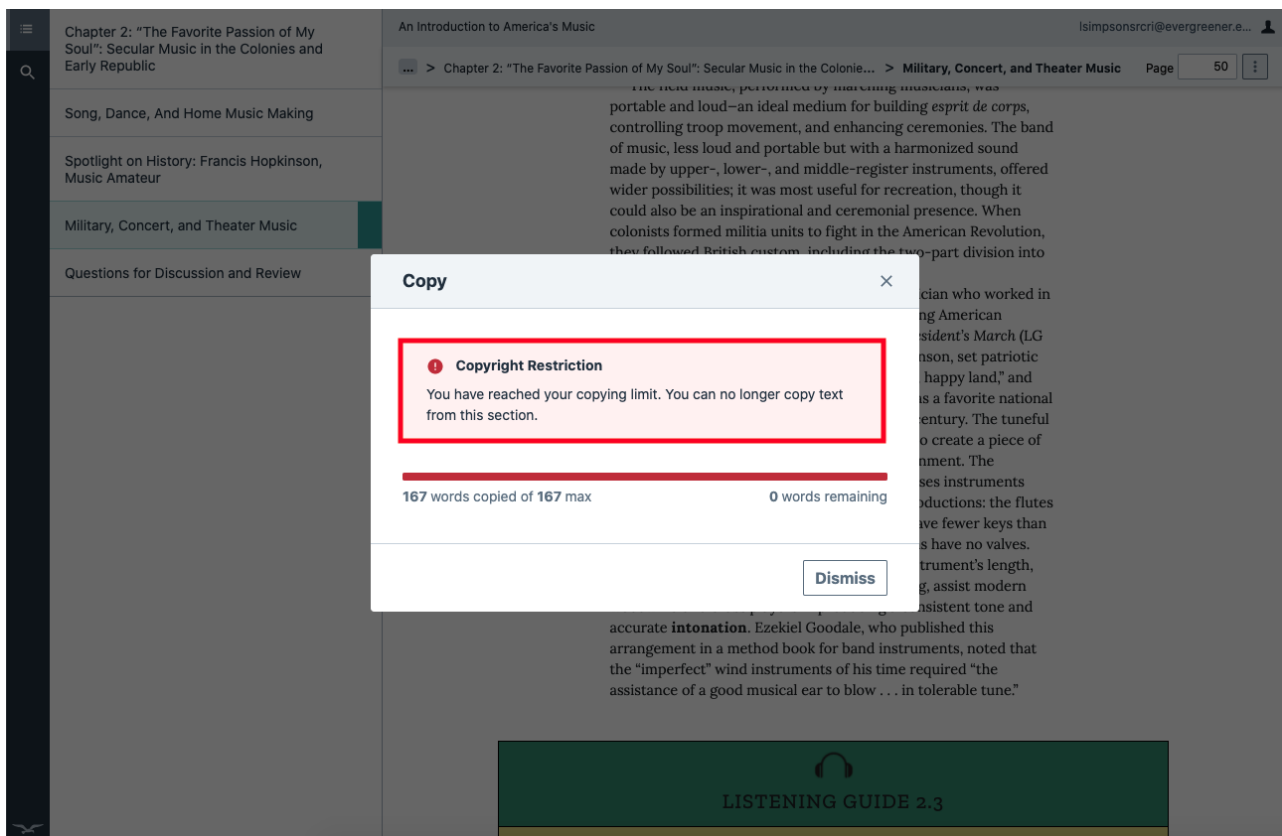
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A screenshot of a document with a redacted section. A red rectangular box covers a portion of the text. Overlaid on this box is a white error message with a red border. The message reads: "Unable to Copy" at the top right, followed by a red exclamation mark icon and the text "Copyright Restriction". Below this, it says "The section of text you have selected cannot be copied due to restrictions on content rights." At the bottom right of the error message is a button labeled "Dismiss". The background text is partially visible and includes phrases like "The discovery of America," "the British writer Adam Smith announced in his celebrated work", "The Wealth of Nations", "the two great and most important", "recorded", "the", "the", "millions of", "the West", "nan", "ances of", "ents a", "had", "a of", "Europe", "Columbus, the interconnection of cultures and peoples has taken place on a global scale. The", "the peoples of the American continents and Europe, previously unaware of each other's existence,", "were thrown into continuous interaction. Crops new to each hemisphere crossed the Atlantic,", "reshaping diets and transforming the natural environment. Because of centuries of lack of", "contact with peoples on other continents, the inhabitants of North and South America had", "developed no immunity to the germs that also accompanied the colonizers. As a result, they", "suffered a series of devastating epidemics, the greatest population catastrophe in human", "history. Within a decade of Columbus's voyage, a fourth continent—Africa—found itself drawn", "into the new transatlantic system of trade and population movement. In Africa, Europeans

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